

## CHAPTER 9 REVIEW

***Stoichiometry*****SECTION 2: Ideal Stoichiometric Calculations**

Please make sure to use correct formulas for the calculations.

1. 

The following equation represents a laboratory preparation for oxygen gas:



How many moles of  $\text{O}_2$  form if 3.0 mol of  $\text{KClO}_3$  are totally consumed?

Show full workings of your answer in this box

2. 

Given the following equation:



Show full workings of your answer in this box

3. 

Water can be made to decompose into its elements by using electricity according to the following equation:



How many grams of  $\text{O}_2$  are produced when 0.033 mol of water decompose?

Show full workings of your answer in this box

4. 

Sodium metal reacts with water to produce NaOH according to the following equation:



How many grams of NaOH are produced if 20.0 g of sodium metal react with excess oxygen?

Show full workings of your answer in this box

# Ideal Stoichiometric Calculations Answers

**M Tight**



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