

Lab 1: Gravimetric Analysis of a Metal Carbonate

BACKGROUND

How do chemists determine the identity of a compound? A large variety of analytical techniques and procedures, ranging from instrumental methods such as spectroscopy and chromatography to more classical processes, such as qualitative and gravimetric analyses, have been created to accomplish that task.

OBJECTIVES

In this experiment, you will determine the identity of a Group 1 metal carbonate (M_2CO_3 , where M is the unknown group 1 metal) using gravimetric analysis.

MATERIALS

- 0.20 M Calcium chloride solution
- Unknown carbonate sample, M_2CO_3
- Distilled water
- Analytical balance
- 400 mL beaker (2)
- Bunsen burner
- Crucible, 15 mL
- Crucible tongs
- Drying oven
- Filter paper
- Buchner funnel + rubber collar
- Filter flask + rubber tubing
- Glass stirring rod
- 250 mL graduated cylinder
- Ring stand + ring
- Spatula
- Clay triangle
- Wash bottle
- Watch glass

PROCEDURE

1. Obtain a clean, dry 15 mL crucible.
2. Set up a Bunsen burner on a ring stand beneath a ring clamp holding a clay triangle. Place the crucible in the clay triangle (see Figure 1).
3. Light the Bunsen burner and brush the bottom of the empty crucible with the burner flame for about 1 minute. Turn off the Bunsen burner and allow the crucible to cool.
4. Using crucible tongs to handle the crucible to avoid leaving fingerprints, measure the mass of the clean, empty crucible to the nearest 0.001 g. Record the mass in the Data Table.
5. While the crucible is still on the balance, add approximately 2 g of the unknown carbonate to the crucible. Record the combined mass of the crucible and the unknown carbonate in the Data Table.
6. Place the crucible on the clay triangle. Light the Bunsen burner again and slowly heat the crucible by brushing the bottom of the crucible with the Bunsen burner flame for 2-3 minutes. Set the crucible to cool on a heat-resistant pad.
7. Weigh the crucible on an analytical balance. Record the mass in the Data Table.

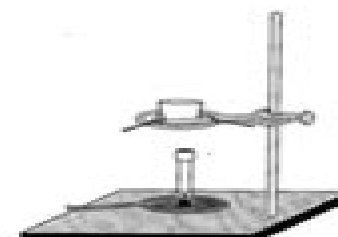


Figure 1.

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