

Equilibrium: Assessment Review

Name _____

SKILL - The student will be able to draw and explain how a catalyst changes a reaction rate graph and discuss other changes that disrupt an equilibrium system.

1. Using the reaction diagram to the right, representing the reaction: $A + B \rightarrow C + D$ draw a new curve that shows how a catalyst would affect the reaction pathway.

2. Place a checkmark next to the actions that would most likely decrease the reaction rate.

- ☐ Lowering the temp of the system
- ☐ Increasing the concentration of substances present
- ☐ Grinding a solid in to fine particles
- ☐ Removing a enzyme catalyst



3. Explain why all reactions have activation energy, using your knowledge of collision model.

4. Explain why different reactions have different activation energies.

5. Describe how the activation energy of a reaction affects the overall rate of the chemical reaction differently than a change in concentration, pressure, or temp.

6. A rule of thumb used by most chemists is that the rate of a chemical reaction can be doubled by increasing the reaction temperature by ten degrees Celsius. Explain this drastic increase in reaction rate using your knowledge of the collision model.

SKILL - The student will be able to write the equilibrium expression, given a reaction.

When a reaction reaches equilibrium the reverse reaction forms reactants at the same rate that the forward reaction produces products. A reaction may appear to stop when it reaches equilibrium because the amount of products and reactants does not change. This process is called dynamic equilibrium.

A reaction that has reached equilibrium can be described with a ratio known as the equilibrium expression. This expression relates the concentration of products divided by the concentration of reactants, raising for the equilibrium constant, K_c . When writing equilibrium expressions, only aqueous and gaseous substances are included.

Balance the following equations & write the equilibrium expressions needed to solve for K_c .



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Table of Contents Equilibrium Unit Webquest And Notes Answer Key

1. Understanding the eBook Equilibrium Unit Webquest And Notes Answer Key
 - The Rise of Digital Reading Equilibrium Unit Webquest And Notes Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Equilibrium Unit Webquest And Notes Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Equilibrium Unit Webquest And Notes Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Equilibrium Unit Webquest And Notes Answer Key
 - Personalized Recommendations
 - Equilibrium Unit Webquest And Notes Answer Key User Reviews and Ratings
 - Equilibrium Unit Webquest And Notes Answer Key and Bestseller Lists
5. Accessing Equilibrium Unit Webquest And Notes Answer Key Free and Paid eBooks
 - Equilibrium Unit Webquest And Notes Answer Key Public Domain eBooks
 - Equilibrium Unit Webquest And Notes Answer Key eBook Subscription Services
 - Equilibrium Unit Webquest And Notes Answer Key Budget-Friendly Options

6. Navigating Equilibrium Unit Webquest And Notes Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Equilibrium Unit Webquest And Notes Answer Key Compatibility with Devices
 - Equilibrium Unit Webquest And Notes Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Equilibrium Unit Webquest And Notes Answer Key
 - Highlighting and Note-Taking Equilibrium Unit Webquest And Notes Answer Key
 - Interactive Elements Equilibrium Unit Webquest And Notes Answer Key
8. Staying Engaged with Equilibrium Unit Webquest And Notes Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Equilibrium Unit Webquest And Notes Answer Key
9. Balancing eBooks and Physical Books Equilibrium Unit Webquest And Notes Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Equilibrium Unit Webquest And Notes Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Equilibrium Unit Webquest And Notes Answer Key
 - Setting Reading Goals Equilibrium Unit Webquest And Notes Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Equilibrium Unit Webquest And Notes Answer Key
 - Fact-Checking eBook Content of Equilibrium Unit Webquest And Notes Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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