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Mitsubishi Lancer 2004 User Manual

Robert R. Redfield



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web this data shows that there was a gain the mass of the lighter which shows that there was a certain amount of butane 0 38 grams 0 02 grams gas that was successfully collected during the gas collection there is a mild influence of barometric pressure during both trials as shown above in the raw data

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web 11 use avogadro's law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a volume of 22.4 l at stp $v_1 n_1 = v_2 n_2$ $0.0896 \text{ l} \cdot n_1 = 22.4 \text{ l} \cdot n_2$ $n_2 = \frac{0.0896 \text{ l} \cdot n_1}{22.4 \text{ l}}$ 12 experimental molar mass g/mole of butane 0.23 g / 0.0040 moles = 58 g/mole 13 the molecular formula of butane is C_4H_{10}

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of the water vapor was found by using the temperature of the water

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