

# Year 11 Chemistry Notes - 2021

- Module 1: Properties and Structure of Matter
- Module 2: Introduction to Quantitative Chemistry
- Module 3: Reactive Chemistry
- Module 4: Drivers of Reactions

## **Module 1: Properties and Structure of Matter** **Mixtures**

A mixture is a combination of two or more pure substances (containing one type of molecule) in which each pure substance retains its individual chemical properties - the mixture itself is impure. There are two kinds of mixtures: heterogeneous and homogeneous.

### **Heterogeneous Mixtures**

Two or more substances intermingle, but remain **physically separate**. Often it is possible to separate the original ingredients by **simple physical means**, such as filtering, centrifuge, decanting or sedimentation.

- Examples include: Dirt + Sand, Oil + Water, Salt + Baking Soda
- A Suspension is a specific type of heterogeneous mixture where particles settle at the bottom

### **Homogeneous Mixtures**

Two or more substances have merged into a **uniform phase**. There are no borders between the substances, but they are not chemically bonded. The **physical properties** of each ingredient can be exploited to separate them.

- Examples include: Saltwater, Copper Sulfate solution
- Saltwater can be distilled (boiled) to separate the water
- The two types include solutions and colloids (particles are present, but are very small and do not settle)

**Physical properties** include magnetism, solubility, density, boiling point, melting point, particle size.

### **Methods of Separation**

Separation Method	Property used	Example
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# Keep It Simple Science Preliminary Chemistry

**SA Dillow**



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