

Worksheet: Writing and Balancing Chemical Reactions

1. Balance the following equations and indicate the type of reaction as formation, decomposition, single replacement, double replacement, hydrocarbon combustion, or other.

- $\text{Cu}_{(s)} + \text{O}_{2(g)} \rightarrow \text{CuO}_{(s)}$
- $\text{H}_2\text{O}_{(g)} \rightarrow \text{H}_{2(g)} + \text{O}_{2(g)}$
- $\text{Fe}_{(s)} + \text{H}_2\text{O}_{(g)} \rightarrow \text{H}_{2(g)} + \text{Fe}_3\text{O}_{4(s)}$
- $\text{AsCl}_{3(aq)} + \text{H}_2\text{S}_{(aq)} \rightarrow \text{As}_2\text{S}_{3(s)} + \text{HCl}_{(aq)}$
- $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}_{(s)} \rightarrow \text{CuSO}_{4(s)} + \text{H}_2\text{O}_{(g)}$
- $\text{Fe}_2\text{O}_{3(s)} + \text{H}_2_{(g)} \rightarrow \text{Fe}_{(s)} + \text{H}_2\text{O}_{(l)}$
- $\text{CaCO}_{3(s)} \rightarrow \text{CaO}_{(s)} + \text{CO}_{2(g)}$
- $\text{Fe}_{(s)} + \text{S}_{8(s)} \rightarrow \text{FeS}_{(s)}$
- $\text{H}_2\text{S}_{(aq)} + \text{KOH}_{(aq)} \rightarrow \text{H}_2\text{O}_{(l)} + \text{K}_2\text{S}_{(aq)}$
- $\text{NaCl}_{(s)} \rightarrow \text{Na}_{(s)} + \text{Cl}_{2(g)}$
- $\text{Al}_{(s)} + \text{H}_2\text{SO}_{4(aq)} \rightarrow \text{H}_{2(g)} + \text{Al}_2(\text{SO}_4)_{3(aq)}$
- $\text{H}_3\text{PO}_{4(aq)} + \text{NH}_4\text{OH}_{(aq)} \rightarrow \text{H}_2\text{O}_{(l)} + (\text{NH}_4)_3\text{PO}_{4(aq)}$
- $\text{C}_3\text{H}_8_{(g)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(l)}$
- $\text{Al}_{(s)} + \text{O}_{2(g)} \rightarrow \text{Al}_2\text{O}_{3(s)}$
- $\text{CH}_4_{(g)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(l)}$
- $\text{K}_2\text{SO}_{4(aq)} + \text{BaCl}_{2(aq)} \rightarrow \text{KCl}_{(aq)} + \text{BaSO}_{4(s)}$
- $\text{C}_5\text{H}_{12(l)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(g)}$
- $\text{Ca}(\text{OH})_{2(aq)} + \text{NH}_4\text{Cl}_{(aq)} \rightarrow \text{NH}_4\text{OH}_{(aq)} + \text{CaCl}_{2(aq)}$
- $\text{V}_2\text{O}_5_{(s)} + \text{Ca}_{(s)} \rightarrow \text{CaO}_{(s)} + \text{V}_{(s)}$
- $\text{Na}_{(s)} + \text{ZnI}_{2(aq)} \rightarrow \text{NaI}_{(aq)} + \text{Zn}_{(s)}$
- $\text{C}_7\text{H}_6\text{O}_{3(l)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(l)}$
- $\text{Ca}_{(s)} + \text{N}_{2(g)} \rightarrow \text{Ca}_3\text{N}_{2(s)}$
- $\text{Fe}_2\text{O}_{3(s)} + \text{H}_2_{(g)} \rightarrow \text{Fe}_{(s)} + \text{H}_2\text{O}_{(l)}$
- $\text{C}_{15}\text{H}_{30(l)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(g)}$
- $\text{BN}_{(s)} + \text{F}_{2(g)} \rightarrow \text{BF}_{3(g)} + \text{N}_{2(g)}$
- $\text{C}_{12}\text{H}_{26(l)} + \text{O}_{2(g)} \rightarrow \text{CO}_{2(g)} + \text{H}_2\text{O}_{(g)}$

2. Predict the product(s) along with the states, indicate the type of reaction, and balance the following chemical reactions.

- A solution of lead (II) nitrate is mixed with a solution of sodium iodide.
- Solid zinc sulfide reacts with oxygen in the air.
- Liquid butane ($\text{C}_4\text{H}_{10(l)}$) is used as a fuel to ignite a lighter.
- Barium hydroxide solution is neutralized by adding hydrochloric acid ($\text{HCl}_{(aq)}$).
- Copper metal is placed in a solution of silver nitrate.
- Sulfur burns in oxygen to make sulfur dioxide gas.
- A solution of aluminum sulfate is mixed with a solution of calcium hydroxide.
- Zinc metal is placed in sulfuric acid ($\text{H}_2\text{SO}_{4(aq)}$).
- Aluminum powder is placed in a container filled with chlorine gas.
- Sucrose undergoes cellular respiration.

Identifying And Balancing Chemical Equations Answers

Chapter 8

JL Elias



Identifying And Balancing Chemical Equations Answers Chapter 8:

Identifying And Balancing Chemical Equations Answers Chapter 8 Book Review: Unveiling the Power of Words

In some sort of driven by information and connectivity, the energy of words has be evident than ever. They have the capacity to inspire, provoke, and ignite change. Such could be the essence of the book **Identifying And Balancing Chemical Equations Answers Chapter 8**, a literary masterpiece that delves deep to the significance of words and their affect our lives. Published by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book is key themes, examine its writing style, and analyze its overall impact on readers.

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