

Key

Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, "PerV-nRT", and the universal gas constant $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}$ to solve the following problems:

If pressure is needed in kPa then convert by multiplying by $101.3 \text{ kPa} / 1 \text{ atm}$ to get
 $R = 8.31 \text{ kPa} \cdot \text{L} / (\text{K} \cdot \text{mole})$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 \text{ atm})(12 \text{ L})}{4 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}}$$

$$T = 204.63 \text{ K}$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of 87°C , how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 \text{ atm} \cdot 31 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 360 \text{ K}}$$

$$n = 1.2586 \text{ mol}$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 400 \text{ K}}{60 \text{ L}}$$

$$P = 1.642 \text{ atm}$$

or $P = 166.29 \text{ kPa}$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of 56°C , what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 327 \text{ K}}{0.09 \text{ atm}}$$

$$V = 2310.93 \text{ L}$$

- 5) If I have 17 moles of gas at a temperature of 67°C , and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 \text{ mol} \cdot 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 390 \text{ K}}{88.89 \text{ L}}$$

$$P = 5.34 \text{ atm}$$

or $P = 540.61 \text{ kPa}$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 \text{ atm} \cdot 25 \text{ L}}{0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}} \cdot 300 \text{ K}}$$

$$n = 0.5075 \text{ mol}$$

Ideal Gas Law Problems Sheet 2 Answers

**Frank Schaffer Publications, Joan
DiStasio**



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A Guide to Teaching in the Active Learning Classroom Paul Baepler, J. D. Walker, D. Christopher Brooks, Kem Saichaie, Christina I. Petersen, 2023-07-03 While Active Learning Classrooms or ALCs offer rich new environments for learning they present many new challenges to faculty because among other things they eliminate the room's central focal point and disrupt the conventional seating plan to which faculty and students have become accustomed. The importance of learning how to use these classrooms well and to capitalize on their special features is paramount. The potential they represent can be realized only when they facilitate improved learning outcomes and engage students in the learning process in a manner different from traditional classrooms and lecture halls. This book provides an introduction to ALCs briefly covering their history and then synthesizing the research on these spaces to provide faculty with empirically based practical guidance on how to use these unfamiliar spaces effectively. Among the questions this book addresses are: How can instructors mitigate the apparent lack of a central focal point in the space? What types of learning activities work well in the ALCs and take advantage of the affordances of the room? How can teachers address familiar classroom management challenges in these unfamiliar spaces? If assessment and rapid feedback are critical in active learning, how do they work in a room filled with circular tables and no central focus point? How do instructors balance group learning with the needs of the larger class? How can students be held accountable when many will necessarily have their backs facing the instructor? How can instructors evaluate the effectiveness of their teaching in these spaces? This book is intended for faculty preparing to teach in or already working in this new classroom environment, for administrators planning to create ALCs or experimenting with provisionally designed rooms, and for faculty developers helping teachers transition to using these new spaces. *Fundamentals of Analytical Chemistry* Douglas A. Skoog, 2004 This text is known for its readability combined with a systematic rigorous approach. Extensive coverage of the principles and practices of quantitative chemistry ensures suitability for chemistry majors. **Knowing Thermodynamics** Nicole Marie Gillespie, 2004 Unique Scientific Puzzles Dr. S. Pancharatnam, 2020-04-06 Born and brought up in a sugar factory village, Pancharatnam turned into a good scholar in leading school and college in Pune, then IIT Bombay. This propelled him into some of the world's best universities: UC Berkeley and Stanford. He managed to get away from the ivory towers of USA and devote to more challenging and rewarding Indian chemical industry with more useful R&D and project engineering. Another success story was his own business of specialty filters for the mech engineering industry with over hundred reputed customers. So here he is with a fully enjoyable career of fifty years with over fifty projects and many publications. Having spent all his life in technical investigation and improvements, he has brought to you vast variety of 500 interesting puzzles from various fields. Most are actually encountered in daily life. Many are truly unique and some quite advanced. Further, over 500 jokes are added for relaxing in between. So go ahead, struggle, laugh, and learn a lot. This small book is highly recommended for students of final years of school, all college students.

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engineering problems The first three chapters introduce the basics of problem solving and Excel fundamentals Beyond that the chapters are largely independent of one another Topics covered include graphing data unit conversions data analysis interpolation and curve fitting solving equations evaluating integrals creating macros and comparing economic alternatives

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