

Name: \_\_\_\_\_

Date: \_\_\_\_\_

## Student Exploration: Ideal Gas Law

**Vocabulary:** atmosphere, Avogadro's law, Boyle's law, Charles's law, dependent variable, directly proportional, Gay-Lussac's law, ideal gas, ideal gas constant, ideal gas law, independent variable, inversely proportional, Kelvin temperature scale, kilopascal, mole, pressure, proportionality, STP, volume

**Prior Knowledge Questions** (Do these BEFORE using the Gizmo.)

- Why is it often necessary to add air to your car tires during the winter? Air expands when heated and contracts when cooled – as ambient temperatures get colder, the tires' inflation pressure is going down.

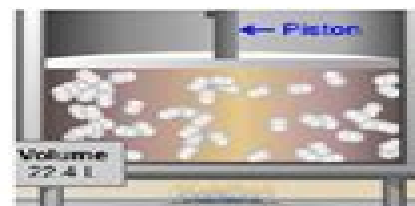


- Why do you think it might be a bad idea to throw an aerosol can into a fire?

Throw an aerosol can into a fire will boil all the liquid contents into gases, which will at that point be highly compressed. Even if the valve ruptures and the contents begin venting, the internal pressure will rapidly build to the point that the can will rupture explosively.

### Gizmo Warm-up

The *Ideal Gas Law* Gizmo shows molecules moving within a chamber fitted with a movable piston. As the piston moves up and down, the **volume** of the chamber changes. Since gases expand to fill their container, any changes in the volume of the chamber changes the volume of the gas within.



- Next to **Dependent variable**, check that **Volume** is selected. Using the green slider, change the **pressure**. Note what happens to the temperature, volume, and amount of gas.

What changes? Volume What stays the same? Temperature temperature and amount of gas

- Using the purple slider on the tank of gas, adjust the number of **moles**, or amount of gas.

What changes? volume What stays the same? Pressure and temperature

# Ideal Gas Law And Answer Key

**Kaplan Test Prep**



## **Ideal Gas Law And Answer Key:**

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