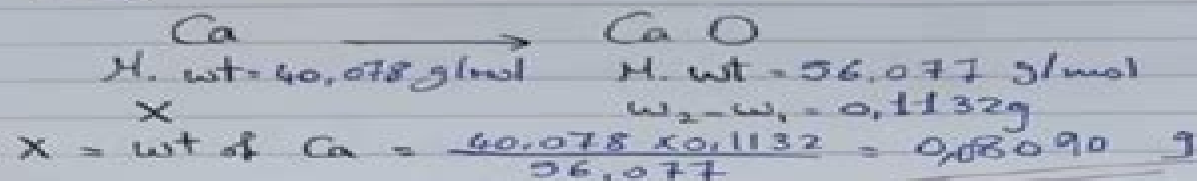


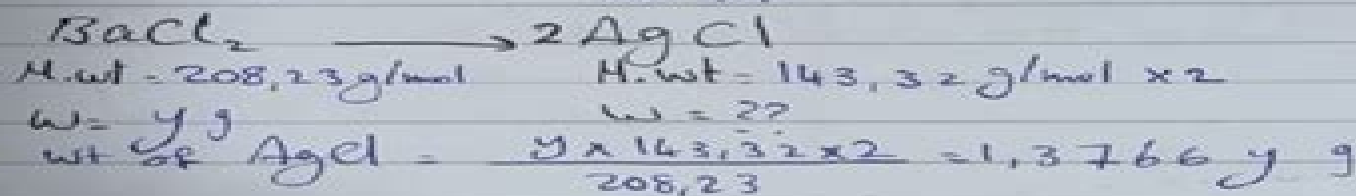
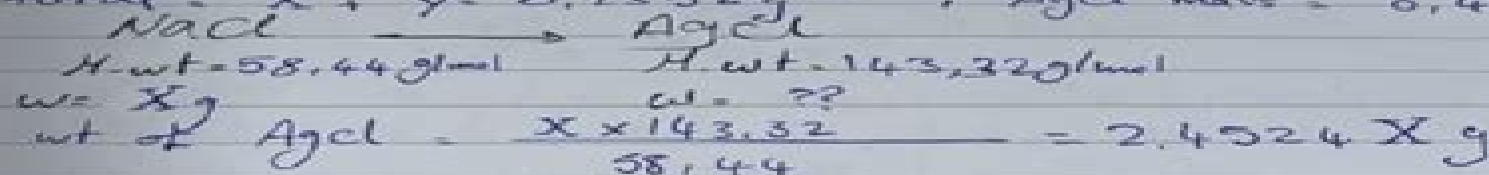
problems (1)

- 1) $w_1 = 26.6002 \text{ g}$, $w_1 = \text{mass of empty crucible}$
 $w_2 = 26.7134 \text{ g}$, $w_2 = \text{mass of crucible and CaO}$
 $w_{\text{of CaO}} = w_2 - w_1 = 26.7134 - 26.6002 = 0.1132 \text{ g}$



$$\begin{array}{l} X \longrightarrow 200 \text{ mL} \\ \text{Strength} \longrightarrow 100 \text{ mL} \\ \therefore \text{Strength of Ca} = \frac{0.08090 \times 100}{200} = 0.04045 \text{ g/100 mL sample} \end{array}$$

- 2) $X = \text{mass of NaCl}$, $y = \text{mass of BaCl}_2$
 $\text{total} = X + y = 0.2356 \text{ g}$, $\text{AgCl mass} = 0.4637 \text{ g}$



$$\begin{aligned} \therefore \text{wt of AgCl} &= 2.4524 X + 1.3766 y \\ &= 2.4524 X + 1.3766 (0.2356 - X) = 0.4637 \text{ g} \\ \therefore X &= \text{mass of NaCl} = 0.1296 \text{ g} \quad \therefore \text{NaCl \%} = 55.01 \% \\ \therefore y &= \text{mass of BaCl}_2 = 0.106 \text{ g} \quad \therefore \text{BaCl}_2 \% = 44.99 \% \end{aligned}$$

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**Azariah Thomas Lincoln, James Henri
Walton**



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