

- 2 -

1. Which of the following processes provide evidence of the particulate nature of matter?

I. Capillarity  
II. Diffusion  
III. Filtration  
IV. Osmosis

(A) I and III only  
(B) I and IV only  
(C) II and III only  
(D) II and IV only

2. 'Mass number' is the number of

(A) neutrons plus protons  
(B) neutrons minus protons  
(C) electrons plus neutrons  
(D) electrons plus protons

3. Which of the following elements has 7 electrons in its outer shell?

(A) Hydrogen  
(B) Oxygen  
(C) Nitrogen  
(D) Chlorine

4. An atom that has 20 electrons when it ionizes has the same electronic configuration as

(A) argon  
(B) calcium  
(C) neon  
(D) sodium

5.  $^{23}_{11}\text{Na}$  and  $^{24}_{11}\text{Na}$  are isotopes because they

(A) conduct electricity  
(B) react vigorously with water  
(C) have the same number of protons  
(D) have the same relative atomic mass

6. Sulfur and oxygen are in the same group of the periodic table because

(A) they can react with each other  
(B) they can form covalent compounds  
(C) they have the same number of electrons in their outer shell  
(D) the atomic number of sulfur is 16 and the relative atomic mass of oxygen is 16

Item 7 refers to the following quantities of atoms.

I. 12 g of carbon  
II. 16 g of sulfur  
III. 23 g of sodium  
IV. 12 g of magnesium

(Relative mass number: C = 12; S = 32; Mg = 24; Na = 23)

7. Which of the following pairs represents equal quantities of atoms?

(A) I and II only  
(B) I and IV only  
(C) II and III only  
(D) II and IV only

8. An ion with a single negative charge may be converted into a neutral atom by

(A) losing a neutron  
(B) gaining a proton  
(C) losing an electron  
(D) gaining an electron

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**Ayd Instone**



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## **Table of Contents Gcse Physics June 2015 Answers**

1. Understanding the eBook Gcse Physics June 2015 Answers
  - The Rise of Digital Reading Gcse Physics June 2015 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Gcse Physics June 2015 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gcse Physics June 2015 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gcse Physics June 2015 Answers
  - Personalized Recommendations
  - Gcse Physics June 2015 Answers User Reviews and Ratings
  - Gcse Physics June 2015 Answers and Bestseller Lists
5. Accessing Gcse Physics June 2015 Answers Free and Paid eBooks
  - Gcse Physics June 2015 Answers Public Domain eBooks
  - Gcse Physics June 2015 Answers eBook Subscription Services
  - Gcse Physics June 2015 Answers Budget-Friendly Options

6. Navigating Gcse Physics June 2015 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Gcse Physics June 2015 Answers Compatibility with Devices
  - Gcse Physics June 2015 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gcse Physics June 2015 Answers
  - Highlighting and Note-Taking Gcse Physics June 2015 Answers
  - Interactive Elements Gcse Physics June 2015 Answers
8. Staying Engaged with Gcse Physics June 2015 Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gcse Physics June 2015 Answers
9. Balancing eBooks and Physical Books Gcse Physics June 2015 Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gcse Physics June 2015 Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Gcse Physics June 2015 Answers
  - Setting Reading Goals Gcse Physics June 2015 Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gcse Physics June 2015 Answers
  - Fact-Checking eBook Content of Gcse Physics June 2015 Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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