

# Year 11 C2 - Discovering Chemistry Topic 1

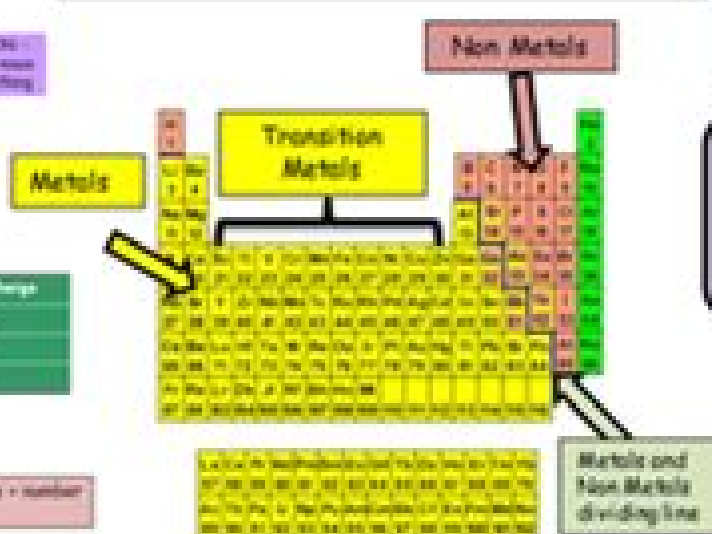


Particle	Relative Mass	Relative Charge
Proton	1	+
Neutron	1	0
Electron	Very small	-

I came up with the first periodic table. I arranged the elements in order of increasing atomic mass

I put elements with similar properties into horizontal rows. I also left gaps incase any new elements were discovered

Dmitri Mendeleev



<sup>14</sup>  
**N**  
7

Electron structure  
+ 2, 5

<sup>24</sup>  
**Mg**  
12

Electron structure  
+ 2, 8, 2

<sup>40</sup>  
**Ca**  
20

Electron structure  
+ 2, 8, 8, 2

**He**

4 ← **MASS NUMBER** = number of protons + number of neutrons

2 ← **PROTON NUMBER** = number of protons (atomic)

CHARGE

## Isotopes

Arrange in a table with a different number of neutrons

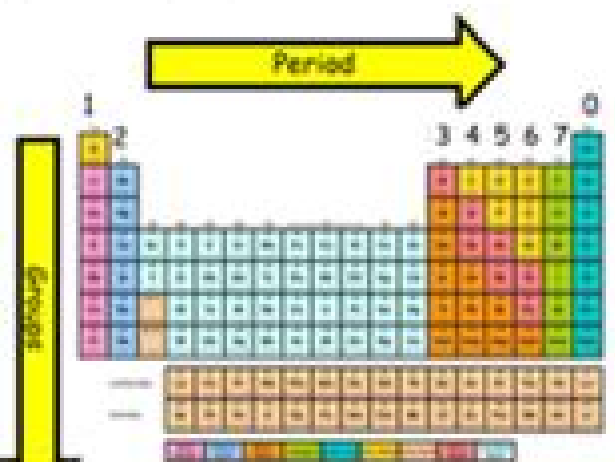
Notice that the mass number is different. How many neutrons does each isotope have?

<sup>16</sup>  
**O**  
8

<sup>17</sup>  
**O**  
8

<sup>18</sup>  
**O**  
8

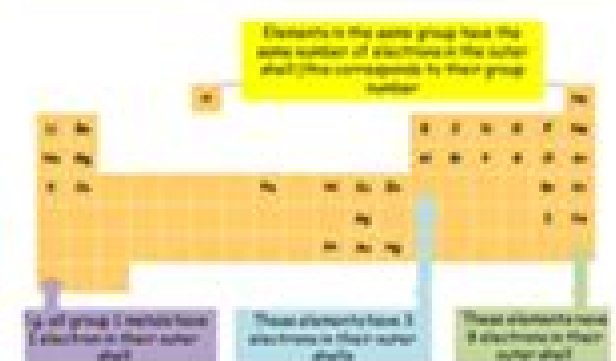
Each isotope has 8 protons - if it didn't then it just wouldn't be oxygen any more



How can an atom have a *decimal* for its mass?

This is because out of every four naturally occurring chlorine atoms, 3 have a mass of 35 and 1 has a mass of 37 so the average atomic mass is:

$35 \times 3 + 1 \times 37 = 106$



(Most important) All of the elements in the same group have similar **PROPERTIES**. This is how I thought of the periodic table in the first place. This is called **PERIODICITY**.

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**SA Adler**



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