Equilibrium Practice Problems

1. Write the equilibrium expression for each of the following reactions:

$N_{2\,(g)} + 3\,H_{2\,(g)} \leftrightarrow 2\,NH_{3\,(g)}$

$$K = \frac{[NH_3]^2}{[N_2]\,[H_2]^3}$$

$I_{2\,(s)} + Cl_{2\,(g)} \leftrightarrow 2\,ICl_{\,(g)}$

$$K = \frac{[ICl]^2}{[Cl_2]}$$

$NO_{2\,(g)} \leftrightarrow NO_{\,(g)} + \frac{1}{2}\,O_{2\,(g)}$

$$K = \frac{[NO]\,[O_2]^{1/2}}{[NO_2]}$$

2. The dissociation of acetic acid, CH_3COOH, has an equilibrium constant at 25°C of 1.8×10^{-5}. The reaction is

$CH_3COOH_{\,(aq)} \leftrightarrow CH_3COO^-_{\,(aq)} + H^+_{\,(aq)}$

If the equilibrium concentration of CH_3COOH is 0.46 moles in 0.500 L of water and that of CH_3COO^- is 8.1×10^{-3} moles in the same 0.500 L, calculate $[H^+]$ for the reaction.

$$K = \frac{[CH_3COO^-]\,[H^+]}{[CH_3COOH]}$$ Plug in known values and solve.

$$1.8 \times 10^{-5} = \frac{[8.1 \times 10^{-3}\ \text{moles}/0.500\ \text{L}]\,[H^+]}{[0.46\ \text{moles}/0.500\ \text{L}]}$$

$$[H^+] = 1.0 \times 10^{-3}\ M$$

3. Indicate the effect of a catalyst, pressure, temperature and concentration on each of the following on:

		catalyst	pressure	temperature	concentration
a.	speed	presence increase	direct relationship	direct relationship	direct relationship
b.	pos^n	nothing	based on moles	based on enthalpy of rxn	inverse relationship

Equilibrium Shift Practice Problems

B Lingard

Equilibrium Shift Practice Problems:

Whispering the Strategies of Language: An Emotional Journey through **Equilibrium Shift Practice Problems**

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