

## **SPECTROPHOTOMETRIC DETERMINATION OF AN EQUILIBRIUM CONSTANT**

### **Lab report**

Name:

ID #:

TA name:

Lab Section:

Lab Date:

1. Looking at Table 5 (data sheet), how is  $K_c$  related to the initial concentrations (  $[Fe^{3+}]_0$  and  $[SCN^-]_0$  )? Is it what you expected? (Explain briefly) (5 marks)

Looking at Table 5 in the results, it is somewhat evident that  $K_c$  is actually independent of the initial concentrations of the reactants since they are relatively consistent among the five test tubes. Furthermore, the  $K_c$  values are much more relative as opposed to the significant changes in  $K_c$  due to temperature variation which holds true because the equilibrium constant is affected by temperature.

However, looking at the data for any trends, it appears that  $K_c$  drops as the concentration of  $[SCN^-]$  increases, but this should not be true. So, accounting for relative human error and limited machine capabilities, the value only fluctuated minimally from the average which was 204 (M<sup>-1</sup>). Thus the results are not exactly as I had expected, but ignoring the small deviations, the results are based on Le Chatelier's Principle which states that when a chemical system in equilibrium is disturbed there will be a shift in the system to reduce the effect of that disturbance. So, by increasing  $[SCN^-]$  from tube 1 to 5, there should be no change of  $K_c$  since equilibrium will shift to the right of the reaction to increase the concentration of  $[FeSCN^{2+}]$  and balance out.

$$K_c = \frac{[FeSCN^{2+}]_{eq}}{[Fe^{3+}] * [SCN^-]}$$

2. To obtain  $[FeSCN_{2-}]_{eq}$  in tube 5, you made the assumption that 100% of the ions  $SCN^-$  had reacted. Now that you know the value for  $K_c$  (average in table 5), build an ICE diagram to calculate the true percentage of  $SCN^-$  reacted in tube 5. Comment on the validity of the assumption. (Make sure you include the ICE diagram in your answer) (10 marks)

ICE Table

	$[Fe^{3+}]$	$+ [SCN^-]$	$\rightleftharpoons [FeSCN_{2-}]$
Initial	0.180 M	0.000199 M	0 M
Change	-x	-x	+x
Equilibrium	$(0.180 - x)$ M	$(0.000199 - x)$ M	x M

# Equilibrium Constant Lab Report Discussion

**JE Gale**



## Equilibrium Constant Lab Report Discussion:

Questions & Answers About Block Scheduling John Brucato, Donald Gainey, 2014-04-11 For administrators and others involved in the transition to block schedules this book provides answers to the complex and challenging questions raised by the curious and the skeptical It demonstrates how to overcome obstacles to systemic school improvements Scientific and Technical Aerospace Reports ,1992 U.S. Government Research Reports ,1964 **Abstracts of Papers** American Chemical Society. Meeting, 1986 Nuclear Science Abstracts ,1965 **Energy Research Abstracts** ,1989 **Technical Abstract Bulletin** ,1965 *The Theory of Kinetics* R.G. Compton, C.H. Bamford, C.F.H. Tippet†, 2012-12-02 The Theory of Kinetics covers the practice and theory of kinetics and the kinetics of inorganic and organic reactions in gaseous and condensed phases and at interfaces This text is composed of five chapters and starts with a review of the kinetic characterization of complex reaction systems The succeeding chapter describes the formal and radical kinetics as well as the energy factor in chain reactions These topics are followed by a survey of the theory of the kinetics of elementary gas phase reactions and the unimolecular reaction of activated chemical species The discussion then shifts to the general properties reactions and the theory of elementary reactions in solution The last chapter examines the theory of kinetics of solid state reactions This book is of great value to physical inorganic and organic chemists **U.S. Government Research & Development Reports** ,1965 **The Naval Medical Research Institute, 1942-1962** Naval Medical Research Institute (U.S.), 1966 *Partners in Innovation* Elaine Seymour, 2005 Seymour argues from evidence that effective deployment adequate professional education and collegial collaboration between faculty and their TAs are critical in ensuring the future quality of science education BOOK JACKET **Bibliography of Scientific and Industrial Reports** ,1965-07 **Chemical Abstracts** ,1910 Yucca Mountain Site Characterization Project Bibliography , 1994-1995 ,1996 **Bibliography of Scientific and Industrial Reports** ,1964-04 **The Addison-Wesley Book of Apple Computer Software 1983** Jeffrey Stanton, Robert P. Wells, Sandra Rochowansky, 1983 **Government Reports Announcements** ,1973-11-10 **The Addison-Wesley Book of Apple Software 1984** Jeffrey Stanton, 1984 The Naval Medical Research Institute, 1942-1962 Naval Medical Research Institute (U.S.), 1966 U.S. Government Research and Development Reports ,1965

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