

Algebra 2 Unit 12: Statistical Notes

Variability and Sampling

Data is everywhere. In its raw form, pages, it's in our science classes, it shows up in economics, statistics and anywhere else that variability occurs. Variability is simply the property of outcomes being different. The tools of statistics are designed to explain this variability. There are many types of variability. It is good to understand these various in order to understand their origin, how we measure variability.

Example 1 The following types of variability can change variability of a data set. For each, give an example that may hold.

- (a) **Observational or Measurement Variability:** Variability that is introduced due to either our measuring instruments and being present enough to distinguish as have two different people read the measurement.

2 students measure time for a ball to roll down a ramp
→ different reaction times

- (b) **Natural Variability or Inherent Individual Variability:** Variability that occurs due to the fact that members of populations are simply different.

2 people work out same amount of time
→ different people use different amounts of weight

- (c) **Random Variability:** This type of variability is the result of noise and error. It can be because our tools assigned our populations or samples to two or more treatment groups and then observing the variability between the groups.

2 groups taking a test, one group allowed to sleep 8 hours, other group only 4 hours → different tests

- (d) **Sample Variability:** This is the type of variability that occurs when we take multiple samples. There is population variability. There is going to still be difference due to the variability of the samples or means.

Multiple samples of people surveyed to determine political candidate preference.

Remember, through all of our work in this book, we are really trying to explain the variability of data within either a population or a sample and then using this to determine if the variability can be attributed to one of the factors above or the variability of the others.

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+ \rightarrow 2\text{AgCl}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver i ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 2ag aq 2f aq 2nh₄ aq cr₂o₇²⁻ aq ag₂cr₂o₇ s 2nh₄ aq 2f aq because two nh₄ aq

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