

THE MOLE WORKSHEET

1. What is a mole of a substance?

A mole of a substance is 6.02×10^{23} entities (atoms or molecules) of a substance. Therefore, one mole of oxygen has 6.02×10^{23} atoms, which will then have a mass of 16.00 g.

2. What does the term molar mass mean?

A molar mass is the mass of one mole of a substance. Oxygen has a molar mass of 16.00 g/mol (for every mole of oxygen there is, it weighs 16.00 g).

3. What is the mass of the following substances?

(a) 3.7 mol of K

$$\begin{aligned} m &= n \times M \\ &= 3.7 \text{ mol} \times 39.10 \text{ g/mol} \\ &= 144.67 \text{ g} \end{aligned}$$

b) 0.0067 mol of SO_2

$$\begin{aligned} m &= n \times M \\ &= 0.0067 \text{ mol} \times 64.06 \text{ g/mol} \\ &= 0.43 \text{ g} \end{aligned}$$

4. Determine the number of moles of substance represented by each of the following masses.

(a) 5.6 g of He

$$\begin{aligned} n &= m / M \\ &= 5.6 \text{ g} / 4.0 \text{ g/mol} \\ &= 1.4 \text{ mol} \end{aligned}$$

(b) 67 g of MgCl_2

$$\begin{aligned} n &= m / M \\ &= 67 \text{ g} / 95.21 \text{ g/mol} \\ &= 0.703 \text{ mol} \end{aligned}$$

5. How many atoms or molecules are present in each of the following? (Hint: Use Avogadro's constant)

(Note: If you are working with an element, you are calculating for the # of atoms. If you are working with a compound, you are calculating for the # of molecules).

a) 2.0 mol Ne

$$\begin{aligned} \# \text{ atoms} &= n_{\text{Ne}} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mol}} \\ &= 2.0 \text{ mol} \times \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mol}} \\ &= 1.2 \times 10^{24} \text{ atoms} \end{aligned}$$

b) 4.5 g H_2O

$$\begin{aligned} \# \text{ molecules} &= m_{\text{H}_2\text{O}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \\ &= m_{\text{H}_2\text{O}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \\ &= \frac{m_{\text{H}_2\text{O}}}{M_{\text{H}_2\text{O}}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \\ &= \frac{4.5 \text{ g}}{18.02 \text{ g/mol}} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \\ &= 1.50 \times 10^{23} \text{ molecules} \end{aligned}$$

c) 0.7 mol CO_2

$$\begin{aligned} \# \text{ molecules} &= n_{\text{CO}_2} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \\ &= 0.7 \text{ mol} \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \\ &= 4.21 \times 10^{23} \text{ molecules} \end{aligned}$$

Introduction To The Mole Article Answers

Joseph Davidovits



Introduction To The Mole Article Answers:

Decoding **Introduction To The Mole Article Answers**: Revealing the Captivating Potential of Verbal Expression

In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Introduction To The Mole Article Answers**," a mesmerizing literary creation penned by a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring impact on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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

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