

NAME: \_\_\_\_\_

CLASS: \_\_\_\_\_

DATE: \_\_\_\_\_

1. Balance the following reaction:  
 $\text{Mg(s)} + \text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$

- |                         |         |                         |         |
|-------------------------|---------|-------------------------|---------|
| <input type="radio"/> A | 1,2,1,2 | <input type="radio"/> B | 1,2,2,2 |
| <input type="radio"/> C | 2,1,1,2 | <input type="radio"/> D | 1,2,1,1 |

2. Using the balanced equation in question 1,  
 $\text{Mg(s)} + \text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$   
 How many grams of HCl are consumed by the reaction of 2.50 moles of magnesium?

- |                         |        |                         |       |
|-------------------------|--------|-------------------------|-------|
| <input type="radio"/> A | 218 g  | <input type="radio"/> B | 7.29  |
| <input type="radio"/> C | 5.00 g | <input type="radio"/> D | 182 g |

3. Using the balanced equation in question 1,  
 $\text{Mg(s)} + \text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$   
 What is the mass in grams of  $\text{H}_2$  gas when 4.0 moles of HCl is added to the reaction?

- |                         |        |                         |       |
|-------------------------|--------|-------------------------|-------|
| <input type="radio"/> A | 0.99 g | <input type="radio"/> B | 4.0 g |
| <input type="radio"/> C | 2.0 g  | <input type="radio"/> D | 6.0 g |

4. Using the balanced equation in question 1,  
 $\text{Mg(s)} + \text{HCl(aq)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$   
 How many moles of HCl are consumed in the production of 7.5 moles of  $\text{MgCl}_2$ ?

- |                         |          |                         |           |
|-------------------------|----------|-------------------------|-----------|
| <input type="radio"/> A | 15 moles | <input type="radio"/> B | 3.8 moles |
| <input type="radio"/> C | 23 moles | <input type="radio"/> D | 7.5 moles |

5. Balance the following reaction :  
 $\text{CaC}_2\text{(s)} + \text{H}_2\text{O(l)} \rightarrow \text{C}_2\text{H}_2\text{(g)} + \text{Ca(OH)}_2\text{(aq)}$

- |                         |         |                         |         |
|-------------------------|---------|-------------------------|---------|
| <input type="radio"/> A | 2,1,1,1 | <input type="radio"/> B | 2,1,2,1 |
| <input type="radio"/> C | 1,2,1,1 | <input type="radio"/> D | 1,2,2,2 |

# Grade 11 Physics Exam Questions Stoichiometry

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*Understanding Physics Using Mathematical Reasoning* Andrzej Sokolowski, 2021-08-20 This book speaks about physics discoveries that intertwine mathematical reasoning modeling and scientific inquiry It offers ways of bringing together the structural domain of mathematics and the content of physics in one coherent inquiry Teaching and learning physics is challenging because students lack the skills to merge these learning paradigms The purpose of this book is not only to improve access to the understanding of natural phenomena but also to inspire new ways of delivering and understanding the complex concepts of physics To sustain physics education in college classrooms authentic training that would help develop high school students skills of transcending function modeling techniques to reason scientifically is needed and this book aspires to offer such training The book draws on current research in developing students mathematical reasoning It identifies areas for advancements and proposes a conceptual framework that is tested in several case studies designed using that framework Modeling Newton's laws using limited case analysis Modeling projectile motion using parametric equations and Enabling covariational reasoning in Einstein formula for the photoelectric effect represent some of these case studies A wealth of conclusions that accompany these case studies drawn from the realities of classroom teaching is to help physics teachers and researchers adopt these ideas in practice *Problems and Problem Solving in Chemistry Education* Georgios Tsapalis, 2021-05-17 Problem solving is central to the teaching and learning of chemistry at secondary tertiary and post tertiary levels of education opening to students and professional chemists alike a whole new world for analysing data looking for patterns and making deductions As an important higher order thinking skill problem solving also constitutes a major research field in science education Relevant education research is an ongoing process with recent developments occurring not only in the area of quantitative computational problems but also in qualitative problem solving The following situations are considered some general others with a focus on specific areas of chemistry quantitative problems qualitative reasoning metacognition and resource activation deconstructing the problem solving process an overview of the working memory hypothesis reasoning with the electron pushing formalism scaffolding organic synthesis skills spectroscopy for structural characterization in organic chemistry enzyme kinetics problem solving in the academic chemistry laboratory chemistry problem solving in context team based active learning technology for molecular representations IR spectra simulation and computational quantum chemistry tools The book concludes with methodological and epistemological issues in problem

solving research and other perspectives in problem solving in chemistry With a foreword by George Bodner *Chemistry Education and Contributions from History and Philosophy of Science* Mansoor Niaz, 2015-12-23 This book explores the relationship between the content of chemistry education and the history and philosophy of science HPS framework that underlies such education It discusses the need to present an image that reflects how chemistry developed and progresses It proposes that chemistry should be taught the way it is practiced by chemists as a human enterprise at the interface of scientific practice and HPS Finally it sets out to convince teachers to go beyond the traditional classroom practice and explore new teaching strategies The importance of HPS has been recognized for the science curriculum since the middle of the 20th century The need for teaching chemistry within a historical context is not difficult to understand as HPS is not far below the surface in any science classroom A review of the literature shows that the traditional chemistry classroom curricula and textbooks while dealing with concepts such as law theory model explanation hypothesis observation evidence and idealization generally ignore elements of the history and philosophy of science This book proposes that the conceptual understanding of chemistry requires knowledge and understanding of the history and philosophy of science Professor Niaz's book is most welcome coming at a time when there is an urgently felt need to upgrade the teaching of science The book is a huge aid for adding to the usual way presenting science as a series of mere facts also the necessary mandate to show how science is done and how science through its history and philosophy is part of the cultural development of humanity Gerald Holton Mallinckrodt Professor of Physics these are valence bond VB and molecular orbital MO theories He emphasizes the importance of having rival theories and interpretations in science and its advancement He further argues that this VB MO rivalry is still alive and together the two conceptual frames serve as the tool kit for thinking and doing chemistry in creative manners The author surveys chemistry textbooks in the light of the how the books preserve or not the balance between the two theories in describing various chemical phenomena This Talmudic approach of conceptual tension is a universal characteristic of any branch of evolving wisdom As such Mansoor's book would be of great utility for chemistry teachers to examine how can they become more effective teachers by recognizing the importance of conceptual tension Sason Shaik Saere K and Louis P Fiedler Chair in Chemistry Director The Lise Meitner Minerva Center for Computational Quantum Chemistry The Hebrew University of Jerusalem ISRAEL **Research in Education** ,1972 *Resources in Education* ,1993

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