

Maximum Marks: 70

Time Allowed: Three hours

(Candidates are allowed additional 15 minutes for only reading the paper.

They must NOT start writing during this time.)

Answer all questions in Section A, Section B and Section C.

Section A consists of objective / very short answer questions.

Section B consists of short answer questions.

Section C consists of long answer questions.

The intended marks for questions or parts of questions are given in brackets [].

SECTION A – 14 MARKS

Question 1

Refer to the outline Map of India provided and answer the following questions on the Map:

- (i) Mark and label the following on the given map of India: [3]
- (a) The coastal plain that receives rainfall from Northeast monsoons.
 - (b) The oldest mountain range of India.
 - (c) The plateau rich in minerals in India.
- (ii) The angular value of the latitude marked L is: [1]
- (a) $37^{\circ}4' N$
 - (b) $68^{\circ}7' N$
 - (c) $8^{\circ}4' N$
 - (d) $23^{\circ}30' N$
- (iii) The riverine port marked P is: [1]
- (a) Kolkata
 - (b) Kochi
 - (c) Vishakhapatnam
 - (d) Chennai

This Paper consists of 6 printed pages and a map.

Geography Question Paper 201

Hongru Du



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web the balanced chemical equation for this reaction is $\text{mg s} + 2\text{hno}_3 \text{ aq} \rightarrow \text{mg no}_3 \text{ 2 aq} + \text{h}_2 \text{ g}$ in this equation the s symbolizes solid magnesium aq represents aqueous nitric acid aq signifies magnesium nitrate in a dissolved state and g represents the gaseous hydrogen produced

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web mar 4 2018 $\text{mg} + 2\text{hno}_3 \rightarrow \text{mg no}_3 \text{ 2} + \text{h}_2$ on the left side there are mg 1 magnesium in elemental form h 2 there are 2 molecules of nitric acid each acid has one hydrogen so two molecules of nitric acid is 2 hydrogens n 2 there are 2 molecules of nitric acid each acid has one nitrogen so two molecules of nitric acid is 2 nitrogens o 6

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and two moles of nitric acid HNO_3 react to form one mole of magnesium nitrate $\text{Mg}(\text{NO}_3)_2$ one mole of water H_2O and one mole of carbon dioxide CO_2 show chemical structure image reactants

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$\text{Mg}(\text{NO}_3)_2 + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + \text{HNO}_3$ chemical equation - Apr 07 2023

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balanced total equation $\text{H}_2\text{SO}_4 + 2\text{LiOH} \rightarrow \text{Li}_2\text{SO}_4 + 2\text{H}_2\text{O}$