13 Al -> "e + 26 Ma 40K -> = e + 40 G balance nuclear reaction equations. aluminum-26 undergoes positron emission potassium-40 undergoes beta decay (electron) 238 () -> 4 He + 234 74 uranium-238 undergoes alpha decay electron configurations. 152 25226 35236 451345 Write the electron configuration for C. Write the electron configuration for Cr. Write the shorthand notation for Bi. Die] 6524545 d 106p3 NOMENCLATURE [naming compounds, writing formulas] Write the formula for each of the following compounds: a. Sulfur hexafluoride SFS b. Lithium nitride 612 10 c. Chromium (III) carbonate Cra (103) d. Tin (II) chloride 50 Ce2 e. Ammonium acetate NHY CZ HOR (OF CHOCOONH) f. Mercury (I) chloride Hazelez g. Potassium bromate K 81-03 h. Hydrosulfuric acid H25000 i. Chloric acid Sulfurous acid 42500 Name each of the following compounds. a Cuso. Copper (II) Sulfate aluminum fluoride b. AIF c. HI d NO narogen monovale e. H₂Se (14) hydroselenic acid f. HNO: nother and narous acre E. HINO:

Hulf-life

Average atomic mass

h. NaHSO.

Soulden

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

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