

CHEMISTRY 1

FINAL EXAM REVIEW

Table of Contents

30	Legal Stuff
40	Writing and Balancing Equations
50	Writing Compound Names and Formulas
60	Moles and Avogadro's Number
70	Stoichiometry
80	Empirical and Molecular Formula
90	The Periodic Table
100	Gas Law Calculations
110	Molarity and Dilution
120	Dimensional Analysis and Unit Conversions
130	Drawing Lewis Structures
140	Phase Diagrams and Phase Transitions
150	Electron Configurations
160	Intermolecular Forces
170	Periodic Trends
180	Significant Figures
190	Scientific Notation
200	Thermochemistry: Calorimetry and Enthalpy
210	Acids, Bases, and pH
220	Electrons and Quantum Numbers
230	Classifying matter, phase changes, and properties
240	Types of Chemical Reactions, Activity Series, and Solubility
25-440	Answers

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Name _____ Section _____ Date _____

Chemistry Review: Writing and Balancing Equations

Directions: Solve each problem and answer each question

1. Balance each equation below by adding the correct coefficients

- $___ \text{CuCl}_2 + ___ \text{NaBr} \rightarrow ___ \text{NaCl} + ___ \text{CuBr}_2$
- $___ \text{HCN} + ___ \text{Mg(OH)}_2 \rightarrow ___ \text{Mg(CN)}_2 + ___ \text{H}_2\text{O}$
- $___ \text{HNO}_3 + ___ (\text{NH}_4)_2\text{SO}_4 \rightarrow ___ \text{NH}_4\text{NO}_3 + ___ \text{H}_2\text{SO}_4$
- $___ \text{ZnCO}_3 + ___ \text{Al}_2\text{O}_3 \rightarrow ___ \text{ZnO} + ___ \text{Al}_2\text{CO}_3$
- $___ \text{C}_4\text{H}_{10} + ___ \text{O}_2 \rightarrow ___ \text{CO}_2 + ___ \text{H}_2\text{O}$

2. Write and balance each equation. Include phase labels

- Aqueous sodium hydroxide reacts with aqueous sulfuric acid to form aqueous sodium sulfate and water
- Solid iron (III) hydroxide and aqueous sodium carbonate are formed from the reaction of solid iron (III) carbonate and aqueous sodium hydroxide
- Solid aluminum reacts with aqueous hydrobromic acid to form aqueous aluminum bromide and hydrogen gas
- Gaseous ethane (dication hexahydride) reacts with oxygen gas to form gaseous carbon dioxide and gaseous water
- Solid calcium phosphite reacts with aqueous copper (II) nitrate to form aqueous calcium nitrate and solid copper (II) phosphite

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21 CHEMISTRY WORKSHEETS

WORKED ANSWER KEY

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 - The Rise of Digital Reading Florida End Of Course Exam Chemistry Practice
 - Advantages of eBooks Over Traditional Books
2. Identifying Florida End Of Course Exam Chemistry Practice
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Florida End Of Course Exam Chemistry Practice
 - User-Friendly Interface
4. Exploring eBook Recommendations from Florida End Of Course Exam Chemistry Practice
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 - Florida End Of Course Exam Chemistry Practice User Reviews and Ratings
 - Florida End Of Course Exam Chemistry Practice and Bestseller Lists

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 - Highlighting and Note-Taking Florida End Of Course Exam Chemistry Practice
 - Interactive Elements Florida End Of Course Exam Chemistry Practice
8. Staying Engaged with Florida End Of Course Exam Chemistry Practice
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Florida End Of Course Exam Chemistry Practice
9. Balancing eBooks and Physical Books Florida End Of Course Exam Chemistry Practice
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Florida End Of Course Exam Chemistry Practice
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Florida End Of Course Exam Chemistry Practice
 - Setting Reading Goals Florida End Of Course Exam Chemistry Practice
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Florida End Of Course Exam Chemistry Practice
 - Fact-Checking eBook Content of Florida End Of Course Exam Chemistry Practice
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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