

Properties of Acids and Bases

Identification and Classification

Introduction

Acids and bases are useful reagents in the chemistry laboratory, and they play an important role in biology and nature. What are acids and bases? What properties can be used to distinguish between them? Let's explore the properties of acids and bases and identify the characteristic features that will allow us to classify substances as acids and bases.

Concepts

- Acids and bases
- Indicators
- pH scale
- Neutralization
- Active metals
- Conductivity

Background

The word *acid* is derived from the Latin verb *acere*, which means “to (be) sour.” The origin of the word itself reveals a characteristic physical property of acids—they taste sour. Lemons, oranges and grapefruits are called *citrus* fruits because they contain citric acid, an acidic compound that gives them their sour taste. Although taste is an interesting property of the food we eat, it is not a property that we will use in the chemistry laboratory to classify compounds as acids or bases. The following properties are typically used to classify compounds as acids or bases.

- **Conductivity.** Some acids and bases ionize completely when dissolved in water. Solutions that contain large numbers of dissolved ions conduct an electric current and are called *electrolytes*. Other weaker acids and bases may ionize only partially when dissolved in water and may conduct electricity only weakly—they are called *weak electrolytes*. Substances that do not produce dissolved ions will not conduct electricity and are called *nonelectrolytes*.
- **Reactions with Metals.** Acids react with so-called *active metals*—reactive metals, such as magnesium and zinc—to produce hydrogen gas and solutions of metal ions. The reactions of different metals with acids is a well-known test used to rank metals from most active to least active. Reaction of a single active metal with a variety of different solutions is one of the best methods to identify acids and compare their relative acidity.

- **Effect on Indicators.** Indicators are organic dyes that change color in acidic or basic solutions. One of the oldest known acid–base indicators is litmus, a natural dye obtained from lichens. Its use was described as early as the 16th century. Litmus paper, prepared by soaking paper in a solution of the dye, is often used as a general test for acids and bases. Phenolphthalein is another indicator that shows a color change as solutions change from acidic to basic. Although these indicators are useful for broadly classifying substances as acids or bases, they are not able to distinguish among different levels of acidity or basicity. By using combinations of different indicators, however, it is possible to obtain a spectrum of color changes over a wide range of acidity levels. Universal indicator and pH paper are two products that use combinations of indicators to rank substances from most acidic to least acidic or most basic to least basic.

- **The pH Scale.** The pH scale is a numerical scale used to describe the relative acidity or basicity of a solution related to the concentration of H_3O^+ ions. The abbreviation pH stands for “power of hydrogen”—a difference of one unit on the pH scale corresponds to a power of ten difference in the concentration of H_3O^+ ions. Pure water contains extremely small but equal concentrations of H_3O^+ and OH^- ions due to self-ionization (Equation 1). Acids and bases are substances that alter the concentrations of H_3O^+ and OH^- ions in solution.



- **Neutralization Reactions.** Acids and bases react with each other to give *neutral* products—solutions that are neither acidic nor basic. The products of the neutralization of an acid and a base are an ionic compound (generally referred to as a salt) and water. The amount of acid that will react with a specific amount of base is governed by stoichiometry. Neutralization reactions are frequently used in the laboratory to determine how much of an acidic or basic compound is present in a substance.

Experiment Overview

The purpose of this experiment is to explore the properties of aqueous solutions and to classify them as acidic, basic or neutral. The results will be used to develop working definitions and to study the pH scale for identifying acids and bases.

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