

It's Elementary — March Madness

March Madness is upon us! Your students are probably already working on their next assignment in brachutology, predicting the field for the NCAA Basketball Tournament. (Although the word doesn't appear in Webster's, there is an article on brachutology on Wikipedia, so it must be real.) Let's apply the principles of brachutology to the properties of the elements. The competition among the 64 elements is still, and there is a wealth of information for students to research in this "elementary" version of March Madness!

Review of Element Properties

Define each of the following properties of the elements and briefly describe any periodic trend in the property across a row or down a column in the periodic table.

- Ionization energy
- Atomic radius

Tournament Rules

Here are the rules for predicting the winners in each round of *It's Elementary—March Madness*. In the event of a tie in the properties of two competing elements in any round of the tournament, the element with the larger atomic mass always wins.

- First round: Research the date of discovery of each element. In each bracket, the element that was discovered earlier in its *life* (see element period) wins and proceeds to the second round. If an element has been known since ancient times, assign it a discovery date of zero.
- Second round: Compare the ionization energy of the elements in each bracket. The element with the higher ionization energy is the winner and advances to the Sweet 16.
- Third round (Sweet 16): Compare the group numbers of the elements—the winner is the element with the larger group number using the International (IUPAC) system (Groups 1–18).
- Fourth round: The element with the larger atomic radius wins this round and earns a trip to the Final Four. Use the atomic covalent radius only, not the van der Waals radius.
- Semifinals (Final Four): Solve the following riddles to determine the two elements that will compete for the championship.
 - This "salt-maker" is also a mineralizer where its white salt is scattered into clouds.
 - Once a sulfide and once for services rendered, the ion of this element is now a life or comatose (place) representative.
- Finals: It's often said that there's no _____ in tennis, but it is the winner in this field!

Connecting to the National Standards

This laboratory activity relates to the following National Science Education Standards (1996):

Designing Concepts and Processes: Grades K–12

Systems, order, and organization.

Content Standards: Grades 9–12

Content Standard B: Physical Science, structure of atoms, structure and properties of matter.

References

- John Emsley, *Nature's Building Blocks*, Oxford University Press, Oxford, England (2001).
WebElements™ Periodic Table, <http://www.webelements.com> (accessed February 2015).
Flinn Scientific Periodic Table of the Elements, Catalog No. AP9829.

Flinn Science March Madness Table Answer Key

Victor M. Corman

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