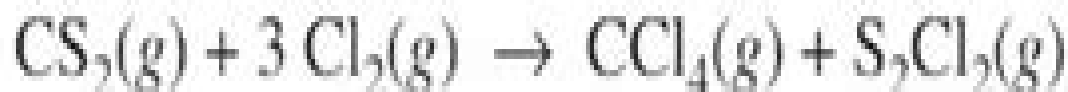


# 2017 AP Chemistry FRQ

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## Question 1



1. Carbon tetrachloride,  $\text{CCl}_4(g)$ , can be synthesized according to the reaction represented above. A chemist runs the reaction at a constant temperature of  $120^\circ\text{C}$  in a rigid 25.0 L container.
  - (a) Chlorine gas,  $\text{Cl}_2(g)$ , is initially present in the container at a pressure of 0.40 atm.
    - (i) How many moles of  $\text{Cl}_2(g)$  are in the container?
    - (ii) How many grams of carbon disulfide,  $\text{CS}_2(g)$ , are needed to react completely with the  $\text{Cl}_2(g)$ ?

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**Thirumalaisamy P.  
Velavancorresponding**



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