

Part 2: Le Chatelier's Principle: Predicting Shifts in Equilibrium

Le Chatelier's Principle:

When a system at equilibrium is stressed, so that it is no longer is at equilibrium, the reaction system shifts in the forward or reverse direction to reestablish equilibrium.

Stresses include:

- Adding or removing a reactant or product from the system
- Changing the temperature (which changed K_{eq} value)

2. Consider the following chemical equilibrium that is endothermic as written:



At equilibrium at a particular temperature, the concentration of $\text{Co(H}_2\text{O)}_6^{2+}$ is slightly greater than the concentration of $[\text{CoCl}_4]^{2-}$. The color of the solution is therefore a light purple.

Predict what will happen when NaCl is added to the solution. Remember that $\text{NaCl}(\text{aq})$ consists of $\text{Na}^+(\text{aq})$ and $\text{Cl}^-(\text{aq})$ ions.

- (a) The reaction will shift to the products
- (b) Concentration of $\text{Co(H}_2\text{O)}_6^{2+}(\text{aq})$ will decrease
- (c) Concentration of $[\text{CoCl}_4]^{2-}(\text{aq})$ will increase
- (d) The color of the solution will become more blue

3. Consider again the following chemical equilibrium that is endothermic as written:



Predict what will happen when the temperature of the equilibrium system is reduced. This is equivalent to removing "heat" from the system.

- (a) The reaction will shift to the reactants
- (b) Concentration of $\text{Co(H}_2\text{O)}_6^{2+}(\text{aq})$ will increase
- (c) Concentration of $[\text{CoCl}_4]^{2-}(\text{aq})$ will decrease
- (d) The color of the solution will become more pink
- (e) Does the value of K_{eq} increase or decrease? Remember that K_{eq} is a ratio of products over reactants at equilibrium.
at equilibrium the constant will remain the same

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