

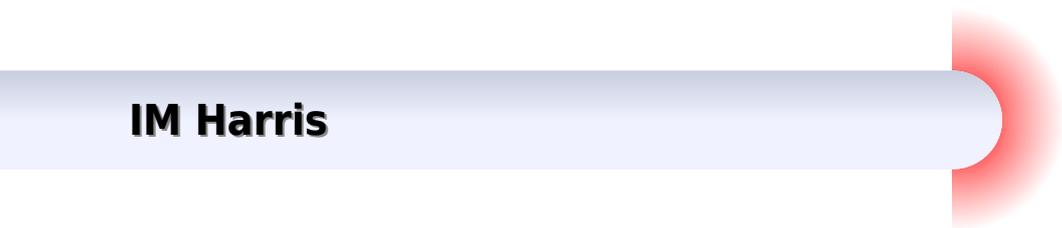
Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The charge on an ion is
a. always positive.
b. always negative.
c. either positive or negative.
d. zero.
- _____ 2. According to the octet rule, a calcium atom has a tendency to
a. lose one electron.
b. lose two electrons.
c. gain one electron.
d. gain two electrons.
- _____ 3. If a compound forms by ionic bonding, which is *not* true?
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.
b. The net charge of the compound is zero.
c. The compound contains just two atoms, each of opposite charge.
d. Several ions group together in a tightly packed structure.
- _____ 4. The only property listed that is *not* characteristic of ionic compounds is
a. high melting point.
b. hardness.
c. lack of crystal structure.
d. good conductor of electricity.
- _____ 5. Which formula listed below represents a polyatomic ion?
a. NH_4^+
b. H_2SO_4
c. NaCl
d. H_2O
- _____ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because
a. ionic substances tend to vaporize at room temperature.
b. ionic substances are brittle.
c. attractive forces between ions are greater than the attractive forces between molecules.
d. the numbers of positive and negative charges are equal in an ionic compound.

Introduction To Chemical Bonding Section Review Answers

IM Harris



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Chemistry Tadashi Okuyama,Howard Maskill,2013-11 Organic Chemistry A mechanistic approach provides readers with a concise review of the essential concepts underpinning the subject It combines a focus on core topics and themes with a mechanistic approach to the explanation of the reactions it describes making it ideal for those looking for a solid understanding of the central themes of organic chemistry Opening with a review of chemical bonding and molecular shape and structure the book then introduces the principal groups of organic compound before exploring the range of reactions they undergo It retains an emphasis throughout on how and why organic compounds behave in the way they do with a chapter on how mechanisms are investigated and the closing chapter describing the principal methods by which the structure and composition of organic compounds are studied With an understanding of organic chemistry being central to the

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