

## Mole Calculation Practice Worksheet

Answer the following questions:

- 1) How many moles are in 25 grams of water?

$$1.4 \text{ moles}$$

$$\frac{25 \text{ g H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \left| \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \right. =$$

- 2) How many grams are in 4.5 moles of  $\text{Li}_2\text{O}$ ?

$$130 \text{ g}$$

$$\frac{4.5 \text{ mol Li}_2\text{O}}{1 \text{ mol Li}_2\text{O}} \left| \frac{29.88 \text{ g Li}_2\text{O}}{1 \text{ mol Li}_2\text{O}} \right. =$$

- 3) How many molecules are in 23 moles of oxygen?

$$1.4 \times 10^{25} \text{ molecules}$$

$$\frac{23 \text{ mol O}_2}{1 \text{ mol O}_2} \left| \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol O}_2} \right. =$$

- 4) How many moles are in  $3.4 \times 10^{23}$  molecules of  $\text{H}_2\text{SO}_4$ ?

$$0.56 \text{ moles}$$

$$\frac{3.4 \times 10^{23} \text{ molecules H}_2\text{SO}_4}{6.02 \times 10^{23} \text{ molecules H}_2\text{SO}_4} \left| \frac{1 \text{ mol H}_2\text{SO}_4}{6.02 \times 10^{23} \text{ molecules H}_2\text{SO}_4} \right. =$$

- 5) How many molecules are in 25 grams of  $\text{NH}_3$ ?

$$8.8 \times 10^{23} \text{ molecules}$$

$$\frac{25 \text{ g NH}_3}{17.03 \text{ g NH}_3} \left| \frac{1 \text{ mol NH}_3}{17.03 \text{ g NH}_3} \right| \frac{6.02 \times 10^{23} \text{ molecules NH}_3}{1 \text{ mol NH}_3} =$$

- 6) How many grams are in  $8.2 \times 10^{22}$  molecules of  $\text{N}_2\text{I}_4$ ?

$$110 \text{ g}$$

$$\frac{8.2 \times 10^{22} \text{ molecules N}_2\text{I}_4}{6.02 \times 10^{23} \text{ molecules N}_2\text{I}_4} \left| \frac{1 \text{ mol N}_2\text{I}_4}{6.02 \times 10^{23} \text{ molecules N}_2\text{I}_4} \right| \frac{789.42 \text{ g N}_2\text{I}_4}{1 \text{ mol N}_2\text{I}_4} =$$

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