



GRAM FORMULA MASS

Name _____

Determine the gram formula mass (the mass of one mole) of each compound below.

1. KMnO_4 $39 + 55 + 16(4) = \boxed{158}$

2. KCl $39 + 35 = \boxed{74}$

3. Na_2SO_4 $23(2) + 32 + 16(4) = \boxed{142}$

4. $\text{Ca(NO}_3)_2$ $40 + (14 + 16 \cdot 3)2 = \boxed{164}$

5. $\text{Al}_2(\text{SO}_4)_3$ $27 + (32 + 16 \cdot 4)3 = \boxed{319}$

6. $(\text{NH}_4)_3\text{PO}_4$ $(14 + 4)(3) + 31 + 16(4) = \boxed{149}$

7. $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ $64 + 32 + 16(4) + 18(5) = \boxed{250}$

8. $\text{Mg}_3(\text{PO}_4)_2$ $24(3) + (31 + 16 \cdot 4)2 = \boxed{262}$

9. $\text{Zn}(\text{C}_4\text{H}_7\text{O}_2)_2 \cdot 2\text{H}_2\text{O}$ $65 + [2(24(4) + 16(2))2 + 18(2)] = \boxed{219}$

Gram Formula Mass Answer Key

Thomas Griffiths

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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