

Chapter 2 Practice Problems

1. What does "X" represent in the following symbol?



- (A) silicon
B) sulfur
C) zinc
D) ruthenium
E) nickel

2. Determine the number of protons, neutrons and electrons in the following:



$$\begin{array}{l} p: 18 \\ n: 22 \\ e: 18 \end{array}$$

3. Determine the number of protons, neutrons and electrons in the following:



$$\begin{array}{l} p: 29 \\ n: 36 \\ e: 29 \end{array}$$

4. What element is defined by the following information?

$$p^+ = 17 \quad n^0 = 20 \quad e^- = 17$$

Cl

5. What species is represented by the following information?

$$p^+ = 12 \quad n^0 = 14 \quad e^- = 10$$



6. Predict the charge that an aluminum ion would have.

$$3A = +3$$

7. Predict the charge that a calcium ion would have.

$$2A = +2$$

8. Calculate the atomic mass of silver if silver has 2 naturally occurring isotopes with the following masses and natural abundances:

Ag-107	106.90509 amu (.518%)	51.84%	= 55.42	> = 108.2
Ag-109	108.90476 amu (.482%)	48.16%	= 52.78	

9. There are 54 electrons, 54 protons, and 78 neutrons in an atom of $^{132}_{54}\text{Xe}$.

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