

5th Grade Science SOLs

Study Guides & Review Worksheets

Sources of Energy

Nonrenewable Energy

Nonrenewable energy resources are natural resources that cannot be replaced after they are used because they take millions of years to form.

Examples

Fossil fuels are formed from the remains of ancient plants and animals.

Fossil fuels include:



Petroleum (oil)



Coal



Natural Gas

Air Pollution



Water Pollution



Conservation

Conservation is the careful use and protection of energy resources. It means using resources in a way that conserves them.

- Turn off lights and electronics when leaving a room.
- Use energy-efficient light bulbs.
- Use energy-efficient appliances.
- Use energy-efficient devices.
- Use natural light.

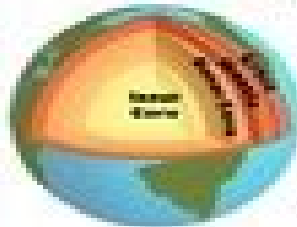
- Take short showers.
- Recycle paper and plastic.
- Recycle food scraps.
- Recycle glass.
- Recycle metal.

Geology

Terms to Know

- Erosion:** The wearing down and changing of rocks and minerals due to natural processes.
- Strata:** The process of soil, rocks, and other sediments being carried away or moved by natural forces, such as wind, water, and ice.
- Deposition:** When lots of soil, rocks, and other sediments settle in a new place.

Layers of the Earth



The Earth has layers. Pressure and temperature increase the deeper you go below the Earth's surface.

- Crust:** The thin, outermost layer of the Earth. It is composed of solid rock.
- Mantle:** The largest layer that is between the crust and the outer core. It is made of hot, semi-solid rock.
- Outer Core:** The liquid layer surrounding the inner core.
- Inner Core:** A solid sphere at the center of the Earth. It is made of materials such as iron and nickel.

Earth's Movements

Divergent Boundaries

- Plates move apart from each other.
- Results in volcanoes and magma formation.



Convergent Boundaries

- Plates move toward each other.
- Results in earthquakes and magma formation.



Transform Boundaries

- Plates slide past each other.
- Results in earthquakes.



Thermal Energy

Earth's thermal energy causes movement within the Earth. It causes currents in the mantle that move the Earth's tectonic plates. This plate movement can result in earthquakes, volcanoes, and magma formation.



Grade 5 Science Study Guide

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a ph regulating agent notice that calcium hydroxide is divalent and thus has twice

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $\text{Ca(OH)}_2 \rightarrow \text{Ca}^{2+} + 2\text{OH}^-$ $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ $2 \times 10^{-2} \text{ M}$ hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web the chemical equation used was $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ the solubility product was determined to be ksp Ca(OH)_2 from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as Ca_3PO_4 is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $\text{Ca}_3\text{PO}_4 \rightleftharpoons 3\text{Ca}^{2+} + \text{PO}_4^{3-}$ $K_{\text{sp}} \text{Ca}_3\text{PO}_4$

post lab report 8 determining the ksp of calcium hydroxide - Jul 06 2023

web lab report 8 determining of calcium hydroxide the purpose of this was to titrate a saturated Ca(OH)_2 solution with a standard hcl solution and find the ksp or the oh concentration was found and then the concentration was found and u for this experiment a titration of calcium hydroxide with hydrochloric acid

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web question determining the ksp of calciumhydroxide $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ for this experiment the titration of

calciumhydroxide with hydrochloric acid wasinvestigated determining the ksp of calcium hydroxide $\text{Ca(OH)}_2 \text{ s} \rightleftharpoons \text{Ca}^{2+} \text{ aq} + 2\text{OH}^- \text{ aq}$ for this experiment the titration of calcium

solved chem 182 experiment 8 determining the ksp of calcium chegg - Sep 27 2022

web question chem 182 experiment 8 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water solution of Ca(OH)_2 is represented in equation form lution of Ca(OH)_2 on s nronic solid that is sparingly soluble in water a saturated aqueous as shown below

determination of the ksp of calcium hydroxide autosaved - Jun 05 2023

web therefore it was determined through error calculations and ease of use that the most reliable valid method to use to get the closest ksp value to the theoretical ksp for Ca(OH)_2 was the determination through ph method 2 introduction calcium hydroxide is a strong base soluble in water

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web plasma membrane homework symbiosis practice worksheet benzoin condensation grade a determining the ksp of calcium hydroxide after reading the lab complete items a b c and d title purpose chemicals and equipment and summary of procedure as described on page 10 of exp 1 on an 8 1 2 x 11 sheet of paper

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web question lab 9 determination of the solubility product ksp of calcium hydroxide introduction ionic compounds that are classified as insoluble based on solubility rules are actually slightly solubla each of these insoluble compounds actually dissolves to

experimental determination of the solubility product of calcium hydroxide - Nov 29 2022

web we will filter the solid calcium hydroxide from the solution and then titrate the dissolved calcium hydroxide in solution with a known concentration of hydrochloric acid using titrimetric calculations we will determine the Ca^{2+} and OH^- concentrations in solution

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web calculate the ksp for calcium hydroxide $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$ ksp 5×10^{-6} find the accepted value of the ksp for calcium hydroxide and compare it with your value discuss the discrepancy and suggest possible sources of experimental error accepted value for ksp for sodium hydroxide is 5×10^{-6}