

Unit 8 Worksheet 3: Molar Concentration

Part A: Describing Molar Concentration

1. Sodium chloride was dissolved in water to produce a 1.5M solution. Explain what this concentration tells us about the NaCl solution. How might a chemist use this ratio?
2. A 45.3 g sample of potassium nitrate is dissolved in enough water to make 225 mL of solution. Determine the molar concentration of the potassium nitrate.
3. Find the molarity of a solution made from 275 g of CuSO_4 dissolved in enough water to make 4.25 L.
4. An alcoholic iodine solution ("tincture" of iodine) is prepared by dissolving 5.15 g of iodine crystals in enough alcohol to make a volume of 225 mL. Calculate the molarity of iodine in the solution.
5. What final volume would be needed in order to prepare a 0.25 M NaCl solution from 5.2 g of NaCl (s)?

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RJ Shavelson



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