

Chemistry 161 - EXAM REVIEW Handout (Lecture 13)

1. Aluminum (Al) is mined from aluminum oxide (Al_2O_3), determine the mass of pure Al that could be extracted from a 467 g sample of Al_2O_3 .

mass% of Al in Al_2O_3 ?

$$\text{Al}_2\text{O}_3 = 101.96 \text{ g/mol}$$

$$26.98 \times 2 = 53.96 \text{ g/mol (mass of Al in } \text{Al}_2\text{O}_3)$$

$$\text{mass\%} = \frac{53.96 \text{ g}}{101.96 \text{ g}} \times 100 = 52.9\%$$



A sample of Ag_2S is found to contain 2.64×10^{24} atoms of Ag. What is the mass of the sample of Ag_2S ?

General Chemistry 1 Exam 4

R Bogdan



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