

Percent Composition and Molecular Formula Worksheet

1. Calculate the percent of nitrogen in urea, NH_2CONH_2 .
2. Calculate the percentage of water in zinc sulfate heptahydrate.
3. Calculate the percentage of oxygen in potassium chlorate.
4. Calculate the percentage of tin in potassium stannate trihydrate, $\text{K}_2\text{SnO}_3 \cdot 2\text{H}_2\text{O}$

Write the molecular formulas of the following compounds:

5. A compound with an empirical formula of C_2OH_4 and a molar mass of 88 grams per mole.
6. A compound with an empirical formula of $\text{C}_4\text{H}_4\text{O}$ and a molar mass of 136 grams per mole.
7. A compound with an empirical formula of CFBrO and a molar mass of 254.7 grams per mole.
8. A compound with an empirical formula of $\text{C}_2\text{H}_5\text{N}$ and a molar mass of 46 grams per mole.

Answer the following questions:

9. The percentage composition of acetic acid is found to be 39.9% C, 6.7% H, and 53.4% O. Determine the empirical formula of acetic acid.
10. The molar mass for question #9 was determined by experiment to be 60.0 g/mol. What is the molecular formula?
11. What's the empirical formula of a molecule containing 18.7% lithium, 16.3% carbon, and 65.0% oxygen?
12. A 50.51 g sample of a compound made from phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39 g of phosphorus atoms were produced. What is the empirical formula of the compound?
13. When 2.5000 g of an oxide of mercury, (Hg_xO_y) is decomposed into the elements by heating, 2.405 g of mercury are produced. Calculate the empirical formula.
14. The compound benzamide has the following percent composition. What is the empirical formula?
C = 69.40 % H= 5.825 % O = 13.21 % N= 11.57 %
15. A component of protein called serine has an approximate molar mass of 100 g/mole. If the percent composition is as follows, what is the empirical and molecular formula of serine?
C = 34.95 % H= 6.844 % O = 46.56 % N= 13.59 %
16. Aniline, a starting material for urethane plastic foams, consists of C, H, and N. Combustion of such compounds yields CO_2 , H_2O , and N_2 as products. If the combustion of 9.71 g of aniline yields 6.63 g H_2O and 1.46 g N_2 , what is its empirical formula?
17. The molar mass of aniline is 93 g/mol. What is its molecular formula?

Formula Writing Cavalcade

ML Morrison



Formula Writing Cavalcade:

This book delves into Formula Writing Cavalcade. Formula Writing Cavalcade is a crucial topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Formula Writing Cavalcade, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Formula Writing Cavalcade
- Chapter 2: Essential Elements of Formula Writing Cavalcade
- Chapter 3: Formula Writing Cavalcade in Everyday Life
- Chapter 4: Formula Writing Cavalcade in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Formula Writing Cavalcade. This chapter will explore what Formula Writing Cavalcade is, why Formula Writing Cavalcade is vital, and how to effectively learn about Formula Writing Cavalcade.
3. In chapter 2, this book will delve into the foundational concepts of Formula Writing Cavalcade. The second chapter will elucidate the essential principles that need to be understood to grasp Formula Writing Cavalcade in its entirety.
4. In chapter 3, the author will examine the practical applications of Formula Writing Cavalcade in daily life. The third chapter will showcase real-world examples of how Formula Writing Cavalcade can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Formula Writing Cavalcade in specific contexts. The fourth chapter will explore how Formula Writing Cavalcade is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Formula Writing Cavalcade. This chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Formula Writing Cavalcade.

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