

Bond Energy

What makes a reaction endothermic or exothermic?

Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** (ΔH). In this activity we will look at one way energy changes can be approximated for chemical reactions.

Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol _{rxn})		Reaction	Change in Energy (kJ/mol _{rxn})
A)	$\text{PCl}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 <i>broken</i>	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_3(\text{g})$	- 966.7 <i>formed</i>
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 <i>broken</i>	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 1,297.9 <i>formed</i>
C)	$\text{PF}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 <i>broken</i>	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_3(\text{g})$	- 1,470.4 <i>formed</i>
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 <i>broken</i>	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 <i>formed</i>

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
- Circle the correct word to complete each sentence below.
 - When bonds are (broken/formed) there is a positive energy change.
 - Breaking bonds is (endothermic/exothermic).
 - When bonds are (broken/formed) there is a negative energy change.
 - Forming bonds is (endothermic/exothermic).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.

- How do the changes in energy for the reverse reactions compare?

The changes in energy for these reactions are the same, but opposite in sign.

- Explain your answer to part a considering what you learned from Questions 1–3 about bond breaking and bond formation.

If a given amount of energy is required to break a bond, then the same amount of energy will be released when the bond is formed.

Flinn Scientific Pogil Activity Bond Energy Answers

Daniela Niemeyer



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Table of Contents Flinn Scientific Pogil Activity Bond Energy Answers

1. Understanding the eBook Flinn Scientific Pogil Activity Bond Energy Answers
 - The Rise of Digital Reading Flinn Scientific Pogil Activity Bond Energy Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Flinn Scientific Pogil Activity Bond Energy Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Flinn Scientific Pogil Activity Bond Energy Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Flinn Scientific Pogil Activity Bond Energy Answers

- Personalized Recommendations
 - Flinn Scientific Pogil Activity Bond Energy Answers User Reviews and Ratings
 - Flinn Scientific Pogil Activity Bond Energy Answers and Bestseller Lists
5. Accessing Flinn Scientific Pogil Activity Bond Energy Answers Free and Paid eBooks
 - Flinn Scientific Pogil Activity Bond Energy Answers Public Domain eBooks
 - Flinn Scientific Pogil Activity Bond Energy Answers eBook Subscription Services
 - Flinn Scientific Pogil Activity Bond Energy Answers Budget-Friendly Options
 6. Navigating Flinn Scientific Pogil Activity Bond Energy Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Flinn Scientific Pogil Activity Bond Energy Answers Compatibility with Devices
 - Flinn Scientific Pogil Activity Bond Energy Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Flinn Scientific Pogil Activity Bond Energy Answers
 - Highlighting and Note-Taking Flinn Scientific Pogil Activity Bond Energy Answers
 - Interactive Elements Flinn Scientific Pogil Activity Bond Energy Answers
 8. Staying Engaged with Flinn Scientific Pogil Activity Bond Energy Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Flinn Scientific Pogil Activity Bond Energy Answers
 9. Balancing eBooks and Physical Books Flinn Scientific Pogil Activity Bond Energy Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Flinn Scientific Pogil Activity Bond Energy Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Flinn Scientific Pogil Activity Bond Energy Answers
 - Setting Reading Goals Flinn Scientific Pogil Activity Bond Energy Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Flinn Scientific Pogil Activity Bond Energy Answers

- Fact-Checking eBook Content of Flinn Scientific Pogil Activity Bond Energy Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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