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**Graphical, Numerical, Algebraic**

FOURTH EDITION



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**R Barnett**



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buffers typically consist of an acid base pair with the acid and base differing by the presence or absence of a proton a conjugate acid base pair for instance one of the buffers that maintain the pH of human blood involves carbonic acid  $\text{H}_2\text{CO}_3$  and its conjugate base the bicarbonate ion  $\text{HCO}_3^-$

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