#### জুনিয়র সুল সাটিফিকেট শরীকা, ২০১৪ গলিত (সৃজনশীল + বহুনির্বাচনি অজীকা) [২০১৪ সালের সিলেবাস অনুযায়ী]

বিষয় কোভঃ 1 0 9

সময়—৩ ঘণ্টা

পূৰ্ণমান—৬০ + ৪০ = ১০০

সৃজনশীল প্রশ্ন

মান--৬৫

দ্রষ্টব্য ঃ—ভান পাশের সংখ্যা প্রস্লের পূর্ণমান জ্ঞাপক। প্রতিটি বিভাগ থেকে কমপক্ষে একটি করে প্রস্ল নিয়ে মোট ছয়টি প্রস্লের উত্তর দাও।

#### ক বিভাগ---পাটীগণিত

একটি আয়তাকার বাগানের দৈর্ঘ্য ৪০ মিটার এবং প্রস্থ ৩০ মিটার। বাগানের ডেতর চারিদিকে ৩ মিটার চওড়া একটি রাস্তা আছে। ২৫ সে,মি, দৈর্ঘ্য এবং ১২·৫ সে, মি, প্রস্থবিশিষ্ট ইট দিয়ে রাস্তাটি পাকা করা হলো। ক) বাগানের ক্বেরফল নির্ণয় কর। 2 8 রান্তার ক্ষেত্রফল নির্ণায় কর। (গ) রাজাটি পাকাকরণে প্রয়োজনীয় ইটের সংখ্যা বের কর। বাবুল হোসেন ৮% ক্ষতিতে একটি ছাগল বিক্রয় করলেন। কিন্ত ছাগলটি ৮০০ টাকা বেশি মূল্যে বিক্রয় করলে তাঁর 21 ৮% লাভ হত। পরবর্তীতে তিনি ছাগলটির ক্রয়মূল্যের সমপরিমাণ টাকা বার্ষিক ১০% মুনাকায় ত বছরের জন্যে ব্যাংকে জমা রাবলেন। (क) ৮% সরল মুনাফায় ৮০০ টাকার ৩ বছরের মুনাফা নির্ণয় কর। হাগদানির ক্রমমূল্য নির্ণয় কর। (গ) উক্ত টাকা ব্যাংকে জমা রাখলে সরল মুনাফা ও চক্রবৃদ্ধি মুনাফার পার্থক্য কত হবে? ধ বিভাগ---বীজগণিত  $U = \{x : x$  স্বাভাবিক সংখ্যা এবং  $x < 8\}$ ;  $P = \{1, 3, 5\};$  $Q = \{x : x$  স্বাভাবিক সংখ্যা এবং  $1 < x \le 7\}$  এবং  $R = \{2, 4, 6\}.$ (ক) P<sup>∞</sup> নির্ণয় কর । 2 (♥) (P ∩ Q) ∪ (Q ∪ R) নির্ণয় কর । 2 (গ) প্রমাণ কর যে, (Q∪R)' = Q' ∩ R'. 3x + 2y = 12, 2x + 3y = 13 দুটি সমীকরণ— (ক) (4, 0) বিন্দৃটি প্রথম সমীকরণকে সিদ্ধ করে কি-না ঘাচাই কর। 2 (খ) অপনয়ন পছতিতে সমীকরণছয়ের সমাধান কর। (গ) লেখচিত্র থেকে সমীকরণছয়ের সমাধান নির্ণয় কর।  $x + \frac{1}{x} = 5$  একটি বীজগাণিতিক সমীকরণ যেখানে x > 0. (ক) 3x + 3/x এর মান নির্ণয় কর। (খ) x<sup>4</sup> + 1/x<sup>4</sup> এর মান নির্ণয় কর। (গ) প্রমাণ কর যে,  $x^2 - \frac{1}{x^2} \neq x^3 + \frac{1}{x^3}$ গ বিভাগ—জামিতি ABCD একটি আয়তক্ষেত্র যার AB=3 সে. মি. উহার AC ও BD কর্ণছয় পরস্পরকে O বিন্দৃতে ছেদ করে। ক) উদ্দীপকের তথ্যগুলো চিত্রের সাহায্যে প্রকাশ কর। 2 (박) 전략적 주로 CT, AC = BD 숙각 AO = CO, BO = DO. AB কে একটি বাছর দৈর্ঘ্য এবং AC = 5 সে, মি.-কে একটি কর্ণের দৈর্ঘ্য ধরে একটি রম্বস অছন কর। [অম্বনের চিহ্ন ও বিবরণ আবশ্যক]

## বাকী পৃষ্ঠাগুলি নিচের লিঙ্কে ক্লিক করে ডাউনলোড করে পড়ুন।

[পর পৃষ্ঠা দ্রষ্টব্য]

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web the substances that undergo change are called reactants the new substances are products sometimes during a chemical reaction one type of reactant will be used up before the other reactants this reactant is the limiting reactant using the limiting reactants gizmo you can determine which reactant is limiting in various scenarios

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web jul 27 2023 by identifying the limiting reactant calculating the amount of product formed and determining the excess reactant chemists can optimize reaction conditions and ensure efficient use of resources in this article we have explored the basics of limiting and excess reactants providing a comprehensive analysis of the pogil answer key for this

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web limiting and excess reactants 5 13 fill in the table below with the maximum moles of water that can be produced in each container q u indicate which reactant limits the quantity of water produced this is the limiting reactant also show how much of the other reactant the reactant in excess will be left over

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web the limiting reagent concept allows us to calculate amounts of reactants used and products formed in a complete chemical reaction based on the stoichiometric relationships in a balanced chemical equation so far we have considered these relationships on the basis of masses and moles

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web workshop limiting reagents and chemical equations pogil worksheet prequisite knowledge ability to solve problems on reaction stoichiometry and mole conversions prequisite vocabulary product reactant chemical equation coefficient part i what are limiting reactants the limiting reactant is the reactant in a chemical reaction that

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for determining the limiting reactant and the amount of excess reactant in a chemical reaction

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web 33mol o2 6mol br2 11mol o2 18mol br2 identify the limiting reactant s and excess reactant s the limiting reactant is o 2 since it would yield the least amount of product 18 mol br 2 the excess reactant is c 2 h 3 br 3 since its complete reaction would have yielded up to 54 mol br 2 think about your result

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web step 4 the reactant that produces a smaller amount of product is the limiting reagent mg mg produces less mgo mgo than does o2 o 2 3 98 g mgo vs 25 2 g mgo therefore mg is the limiting reagent in this reaction step 5 the reactant that produces a larger amount of product is the excess reagent

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web container bodies cylinders tires engines max number of completed cars limiting part a 3 10 9 2 2 engines b 50 12 50 5 c 16 16 16 16 d 4 9 16 6 e 20 36 40

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web please save your changes before editing any questions 5 minutes 1 pt viewing model 3 if 4 moles of hydrogen react with 2 moles of oxygen which of these is the excess reactant is leftover at the end and how many moles of it are leftover hydrogen 2 moles leftover hydrogen 1 mole leftover

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web if you start with 14 82 g of ca oh 2 c a o h 2 and 16 35 g of h2so4 h 2 s o 4 a determine the limiting reagent b determine the number of moles of h2o h 2 o produced c determine the number of grams of caso4 c a s o 4 produced d determine the number of grams of excess reagent left 1 make sure the equation is balanced

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