

# CHAPTER 14: COVALENT BONDING: ORBITALS

## PROBLEMS WITH LEWIS STRUCTURE

- Provides no indication of the **shapes** of the molecules
- $O_2$  would be predicted to be diamagnetic, but it is **paramagnetic**

**Paramagnetic** = has unpaired electrons.



The Lewis structure of  $O_2$  has an  $O=O$  double bond with **no unpaired electrons**, i.e. diamagnetic.

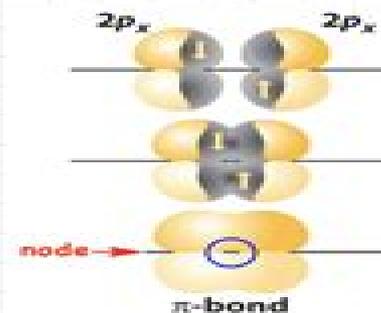
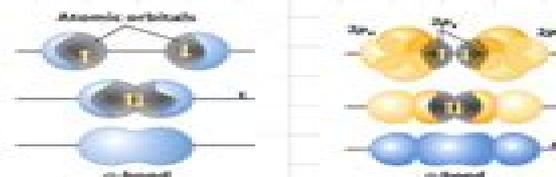
- Many compounds cannot be represented by a single structure
- Resonance structures **differ only in position of electrons** not the nuclei.
- Result in notions of hypervalent compounds and "expansion of octet"
- Can't explain molecules such as  $B_2H_6$  (electron deficient molecules)

## VALENCE-BOND THEORY "LOCALIZED ELECTRON MODEL"

- Covalent bond is created by the overlap of orbitals on two atoms

### TYPES OF OVERLAP: $\sigma$ and $\pi$ Bonds

- **Bonding in  $H_2$** 
  - Sausage shaped orbital, bond called  $\sigma$ -orbital
  - Called  $\sigma$  orbital because it has **no node and is cylindrically symmetrical**
  - Electron density is **concentrated between two nuclei**, but is still highest at the nuclei themselves.
- Bonding in  $N_2$ 
  - $\sigma$  orbital is obtained by overlap of **two  $p_z$ -orbitals**
- Perpendicular p orbitals overlap in a side-by-side manner.
- The resulting orbital has two lobes of electron density and is called a  **$\pi$ -bond**.
- Called  $\pi$ -orbital because it **has a node** and, when viewed down the internuclear axis, it looks like a p-orbital
- $\pi$ -bond has **no electron density between the nuclei along the bond axis**.



# General Chemistry Ii Chapter 14 Covalent Bonding Orbital Notes

**Madan R.L.**



## **General Chemistry II Chapter 14 Covalent Bonding Orbital Notes:**

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general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter WileyPLUS sold separately from text

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Chemical Principles William L. Masterton,Emil J. Slowinski,1969

Modern Physical Organic Chemistry Eric V. Anslyn,Dennis A. Dougherty,2006 Making explicit the connections between physical organic chemistry and critical fields such as organometallic chemistry materials chemistry bioorganic chemistry and biochemistry this book escorts the reader into an area that has been thoroughly updated in recent times

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