

CHAPTER 14: COVALENT BONDING: ORBITALS

PROBLEMS WITH LEWIS STRUCTURE

- Provides no indication of the **shapes** of the molecules
- O_2 would be predicted to be diamagnetic, but it is **paramagnetic**

Paramagnetic = has unpaired electrons.



The Lewis structure of O_2 has an $O=O$ double bond with **no unpaired electrons**, i.e. diamagnetic.

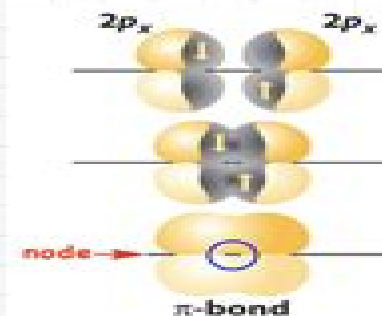
- Many compounds cannot be represented by a single structure
- Resonance structures **differ only in position of electrons** not the nuclei.
- Result in notions of hypervalent compounds and "expansion of octet"
- Can't explain molecules such as B_2H_6 (electron deficient molecules)

VALENCE-BOND THEORY "LOCALIZED ELECTRON MODEL"

- Covalent bond is created by the overlap of orbitals on two atoms

TYPES OF OVERLAP: σ and π Bonds

- **Bonding in H_2**
 - Sausage shaped orbital, bond called **σ -orbital**
 - Called σ orbital because it has **no node and is cylindrically symmetrical**
 - Electron density is **concentrated between two nuclei**, but is still highest at the nuclei themselves.
- Bonding in N_2
 - **σ orbital** is obtained by overlap of **two p_z -orbitals**
- Perpendicular p orbitals overlap in a side-by-side manner.
- The resulting orbital has two lobes of electron density and is called a **π -bond**.
- Called π -orbital because it **has a node** and, when viewed down the internuclear axis, it looks like a p-orbital
- π -bond has **no electron density between the nuclei along the bond axis**.



General Chemistry Ii Chapter 14 Covalent Bonding

Orbital Notes

Madan R.L.



General Chemistry II Chapter 14 Covalent Bonding Orbital Notes:

Chemistry: for B.Sc. Students Semester-I (NEP-UP) Madan R.L.,2022 This textbook has been designed to meet the needs of B Sc First Semester students of Chemistry as per Common Minimum Syllabus prescribed for all Uttar Pradesh State Universities and Colleges under the recommended National Education Policy 2020 Maintaining the traditional approach to the subject this textbook comprehensively covers two papers namely Fundamentals of Chemistry and Quantitative Analysis Important theoretical topics such as molecular polarity weak chemical forces simple bonding theories of molecules periodic properties of atoms basics of organic chemistry mechanism of organic reactions stereochemistry and the mathematical concepts of chemistry are aptly discussed to give an overview of Fundamentals of Chemistry Practical part covering Quantitative Analysis has been presented systematically to help students achieve solid conceptual understanding and learn experimental procedures *EBOOK: GENERAL CHEMISTRY, THE ESSENTIAL CONCEPTS* CHANG,2013-01-07 EBOOK GENERAL CHEMISTRY THE ESSENTIAL CONCEPTS **Chemistry for Degree Students B.Sc. First Year (LPSPE)** Madan R.L.,2022 An outgrowth of more than three decades of classroom teaching experience this book provides a comprehensive treatment of the subject It comprises three parts Inorganic Organic and Physical Chemistry Illustrations and diagrams are provided to help students in understanding the chemical structures and reactions This book will meet the requirements of undergraduate students of B Sc First Year of all Indian universities *General Chemistry* Henry Fuller Holtzclaw,William Harrison Nebergall,William R. Robinson,1984 *General Chemistry with Qualitative Analysis* Ralph H. Petrucci,Robert K. Wismer,1983 *Elements of Physical Chemistry* Peter Atkins,Julio de Paula,2013 Elements of Physical Chemistry has been carefully crafted to help students increase their confidence when using physics and mathematics to answer fundamental questions about the structure of molecules how chemical reactions take place and why materials behave the way they do *Bonding in Electron-Rich Molecules* Richard D. Harcourt,2015-10-30 This second edition was updated to include some of the recent developments such as increased valence structures for 3 electron 3 centre bonding benzene electron conduction and reaction mechanisms spiral chain O4 polymers and recoupled pair bonding The author provides qualitative molecular orbital and valence bond descriptions of the electronic structures for primarily electron rich molecules with strong emphasis given to the valence bond approach that uses increased valence structures He describes how long bond Lewis structures as well as standard Lewis structures are incorporated into increased valence structures for electron rich molecules Increased valence structures involve more electrons in bonding than do their component Lewis structures and are used to provide interpretations for molecular electronic structure bond properties and reactivities Attention is also given to Pauling 3 electron bonds which are usually diatomic components of increased valence structures for electron rich molecules General Chemistry Ralph H. Petrucci,1989 Basic Concepts of Chemistry Leo J. Malone,Theodore O. Dolter,2011-12-27 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address

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General Chemistry William Harrison Nebergall, Frederic Cowles Schmidt, Henry Fuller Holtzclaw, 1968

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Principles of College Chemistry Glenn H. Miller, 1972

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