

UNIT III - THERMOCHEMISTRY

1. Calculate the H-F bond energy given the following data.



$$\Delta H_{\text{rxn}} = -544 \text{ kJ}$$

- A. 1130 kJ/mol B. 565 kJ/mol **C. 283 J/mol** D. 300 kJ/mol E. 134 kJ/mol
2. 1.0 g of the rocket fuel hydrazine (N_2H_4) is burned in a bomb calorimeter. The temperature of the bomb calorimeter changes from 24.63 °C to 28.16 °C. The heat capacity of the bomb calorimeter is 837 J/K. Calculate the heat of combustion per mol of hydrazine in the bomb calorimeter.

A. 95.0 kJ/mol

B. 2.96 kJ/mol

C. 3.55×10^3

D. 92.4 J/mol

E. $1.13 \times 10^5 \text{ kJ/mol}$

3. Calculate the enthalpy change for the following reaction of nitric oxide NO with ozone
- O_3
- .



Given the following data:



- A. -1083.7 kJ B. -693.8 kJ C. -198.8 kJ **D. -93.7 kJ** E. -93.7 kJ
4. An endothermic reaction does no work on the surroundings. What values do you predict for the heat and energy change of this reaction?

A. $q < 0$, $\Delta E > 0$

D. $q > 0$, $\Delta E < 0$

B. $q < 0$, $\Delta E < 0$

E. $q > 0$, $\Delta E = 0$

C. $q > 0$, $\Delta E > 0$

5. M&M candies contain 209.0 Calories per 1.5 oz serving. The fat content accounts for 38.3% of these Calories. What is the energy of the fat content in kJ?

A. 19.1 kJ

B. 874 kJ

C. $1.91 \times 10^{-2} \text{ kJ}$

D. $3.35 \times 10^5 \text{ kJ}$

E. 335 kJ

6. Calculate the heat released when 2.00 g of butane from a cigarette lighter is burned.



A. -90.7 kJ

B. -98.3 kJ

C. -49.2 kJ

D. -108.4 kJ

E. -196.6 kJ

7. When Karl Kaveman adds chilled grog to a granite mug, 10.9 kJ of energy is transferred from the mug to the grog. If the mug has a mass of 625 g and an initial temperature of 25 °C, what is the final temperature of the mug? The specific heat of granite is 0.79 J/g·K.

A. 2.9 °C

B. 14 °C

C. 22 °C

D. 25 °C

E. 47 °C

General Chemistry Exam Questions And Answers

Rachel S Tattersall



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