

09/08/2020

Chapter 1: General Chem Review - Part 1

- 1.1 Reactions: a) molecules collide
b) bonds broken/formed

Organic compounds → compounds w/ carbon atoms

- 1.2 Structural Theory: substances are defined by a specific arrangement
compounds differ based on the way the atoms are bonded

Constitutional isomer: same molecular formula but different structures

Atoms commonly bonded to C: N, O, H & halides

tetravalent
(4 bonds)

trivalent
(3 bonds)

divalent
(2 bonds)

monovalent
(one bond)

- 1.3 Covalent bond: pair of e^- shared by two atoms

Energy released when forming bond → equilibrium internuclear distance

internuclear repulsion if contracted further
(↑ energy)

(bond length)

internuclear distance

energy @ equilibrium
= bond strength

↑ ↑ space = ↓ energy of electrons

Potential Energy & stability: ↓ PE = most stable

What keeps bond @ optimal length?

+ nuclei & - electron → attractive forces

+ nuclei & + nuclei → repulsive forces ← - electron & - electron

Lewis Structure Examples:



- 1.4 Formal charge: valence electrons it has vs. valence electrons need to be neutral
eg. CH_3O



C has 4 valence electrons → no formal charge
Needs 4 more to be neutral

O has 7 electrons → 1- formal charge
Needs 1 less to be neutral

- 1.5 Electronegativity: increase L to R & B to T
F is most EN

↑ carbon usually forms polar covalent bonds

Covalent bond: EN difference < 0.5

Polar covalent bond: EN difference 0.5 - 1.7

Ionic bond: EN difference > 1.7



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Albert A Gayle



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