CHINE SOLL Lab 13 - Caloninetry

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Complete the following density typing into the text box provided. The boxes will increase in size if additional space is necessary. Please he sure to save all your work as an MS Word document to submit properly.

Part 1 Determining the Specific Heat Capacity of Aluminum

A. Disca Table 1.

Complete the Data Table. (7 pts):

ALUMENUM METAL (Trial 1)		
Fre-weighed aluminum metal mass (g)	20.03 g	
Initial temperature of cool water in the calorimeter (oC)	20.0%	
Institut temperature of aluminum metal (name as the boiling water in the pot) (oC)	100.0°C	
Pinal temperature of water/stummum metal in calorimeter (pC)	24 070	
ALUMENUM METAL (Trail 2)		
Fre-weighed sluminum metal mans (g) [Copy this from above since you are using the same metal sample for both trials].	20.03 g	
Initial temperature of cool water in the calorimeter (oC)	18.000	
Initial temperature of aluminum metal (mone as the bosling water in the pot) (o:C)	100 0°C	
Final temperature of waterfalamanum metal in calorimeter (pC)	21170	

It insert the following photo with your name and MEID showing in the photo-

Thoto A. Insert a photo of your calorimeter showing your aluminum metal submerged in the water. The photo should be taken from above. (6 pts)

C. Calculations (Show all calculations with units on each number and the final answer rounded to the correct significant figures for full credit.)

 Calculate the mass of the water by using the water's density (assume it is 1 00 g/mL) and the volume of the water. Remember, density = mass/volume. (4 pts)

$$D = \frac{M}{V} \rightarrow M = D = V \rightarrow M = \left[1.00 \frac{Q}{ml}\right](75 mL) = 75 \text{ galarminum (Trial 1)}$$

$$D = \frac{M}{V} \rightarrow M = D \times V \rightarrow M = \left[1.00 \frac{\alpha}{ml}\right](75 mL) = 75 \text{ gaAlaminum (Trud 2)}.$$

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Padhraic Smyth

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