

1. Which electron configuration represents an atom in an excited state?
 A) $1s^2 2s^2 2p^6 3p^1$ C) $1s^2 2s^2 2p^6 3s^2 3p^2$
 B) $1s^2 2s^2 2p^6 3s^2 3p^1$ D) $1s^2 2s^2 2p^6 3s^2$
2. All of the elements in Period 3 have a total of 2 electrons in the
 A) 2s sublevel C) 2p sublevel
 B) 3s sublevel D) 3p sublevel
3. Which atom in the ground state has three unpaired electrons in its outermost principal energy level?
 A) Li C) N
 B) B D) Ne
4. What is the total number of valence electrons in an atom with the electron configuration $1s^2 2s^2 2p^6 3s^2 3p^3$?
 A) 6 C) 3
 B) 2 D) 5
5. Which orbital notation represents an atom of beryllium in the ground state?
 A)

1s	2s	2p
↑	↑	↑

 B)

↑	↑	↑
↑	↑	↑

 C)

↑	↑	↑
↑	↑	↑

 D)

↑	↑	↑
↑	↑	↑
6. Which element has a completely filled third principal energy level?
 A) Ar C) Fe
 B) N D) Zn
7. Which atom in the ground state has only 3 electrons in the 3p sublevel?
 A) phosphorus C) argon
 B) potassium D) aluminum
8. What is the total number of occupied principal energy levels in a neutral atom of neon in the ground state?
 A) 1 C) 3
 B) 2 D) 4

9. Which sublevel configuration correctly represents a completely filled third principal energy level?
 A) $3s^2 3p^5 3d^6$ C) $3s^2 3p^6 3d^{10}$
 B) $3s^2 3p^2 3d^{10}$ D) $3s^2 3p^6 3d^6$
10. Which atom in the ground state contains a partially filled 3p orbital?
 A) argon C) potassium
 B) calcium D) aluminum
11. In the ground state, the atoms of elements in Period 2 all have the same number of
 A) protons C) 1s electrons
 B) neutrons D) oxidation states
12. Which is the electron configuration for a neutral atom in the ground state?
 A) $1s^2 2s^2 3s^1$ C) $1s^2 2s^2 2p^6 3p^1$
 B) $1s^2 2s^2 2p^4 3s^1$ D) $1s^2 2s^2 2p^6 3s^1$
13. Which represents the electron configuration of an isotope of oxygen in the ground state?
 A) $1s^2 2s^2 2p^1$ C) $1s^2 2s^2 2p^2$
 B) $1s^2 2s^2 2p^2$ D) $1s^2 2s^2 2p^4$
14. The total number of sublevels in the fourth principal energy level of an atom is
 A) 1 C) 3
 B) 2 D) 4
15. Which is a possible electron configuration for argon in the excited state?
 A) $1s^2 2s^2 2p^7 3s^2 3p^3$ C) $1s^2 2s^2 2p^6 3s^1 3p^2$
 B) $1s^2 2s^2 2p^5 3s^2 3p^6$ D) $1s^2 2s^2 2p^6 3s^2 3p^5 4s^1$
16. Which is a correct description of the shape and spatial orientation of the p orbitals in an atom?
 A) All have the same shape but a different orientation.
 B) All have the same shape and the same orientation.
 C) All have a different shape and a different orientation.
 D) All have a different shape but the same orientation.
17. A maximum of 6 electrons can occupy
 A) an s orbital C) a p orbital
 B) an s sublevel D) a p sublevel
18. How does the ground state electron configuration of the hydrogen atom differ from that of a ground state helium atom?
 A) Hydrogen has one electron in a higher energy level.
 B) Hydrogen has two electrons in a lower energy level.
 C) Hydrogen contains a half-filled orbital.
 D) Hydrogen contains a completely filled orbital.

Electron Distribution Review Worksheet Answers

D Keegan



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Development of New Tools for Local Electron Distribution Analysis Eloy Ramos Cordoba, 2014 *Some Factors Affecting the Electron Distribution in Atoms and Molecules* Richard J. S. Crossley, 1964 *On the Electron Distribution in Atoms and Ions* Jarl Axel Wasastjerna, 1933 *Relativistic Two-Temperature Electron Distributions* Howard E. Brandt, HARRY DIAMOND LABS ADELPHI MD., 1980 An electron distribution function is formulated to appropriately describe a collisionless isotropic relativistic and nonequilibrium electron gas consisting of two populations of electrons one in temporal equilibrium at some cold temperature and the other at a hot temperature Characteristics of this two temperature relativistic electron distribution are discussed Author **Electron Distribution Function and Mobility in III-V Compounds for High and Low Fields** S. H. Soh, B. W. Abraham-Shrauner, WASHINGTON UNIV ST LOUIS MO DEPT OF ELECTRICAL ENGINEERING., 1984 Electron distributions are calculated for low and high electric fields in III V semiconductors with polar optical scattering Approximate analytic expressions are derived by two analytic methods and compared to results found by a perturbation method and iteration method for low electric fields Electron mobility is calculated For high electric fields an expansion inversely in the electric field gives an approximate electron distribution function valid at low energies A non parabolic energy band and p wave scattering they and interdependent do not affect the asymmetrical part of the distribution function within the central valley hence they do not cause the negative differential resistance **Basic Aspects of Secondary Electron Distributions**, 1975 **The Electron Distribution in the Bound Water Molecule** Kersti Hermansson, 1984 Electron Distribution Function in a Nonuniform, Magnetized, Weakly Photoionized Gas Application to a Model Ionosphere, 1975 The author presents a plasma kinetic model for the one particle electron distribution function in a dilute nonuniform magnetized weakly photoionized gas The model includes a phenomenological treatment of photoionization of the neutral species by an externally produced photon flux electron ion recombination electron neutral particle attachment elastic excitation deexcitation and ionizing electron neutral particle collisions and elastic electron electron and electron ion collisions Using known methods the author obtains approximate analytic formulae for the steady state electron distribution function at energies much larger than the electron thermal energy For a model of the earth's ionospheric E and F regions at low latitudes numerical solutions for the steady state electron distribution function are given and compared to a curve fitted distribution function of Maxwellian form It is shown that the effective rate of production of electrons per unit volume is significantly enhanced due to ionizing electron neutral particle collisions Results for the electron concentration and the electron temperature are also given On the Validity of the Two-term Approximation of the Electron Distribution Function Robert Edward Robson, Kailash Kumar, 1971* **Electron Distributions and the Chemical Bond** P Coppens (ed), 1982 **Electron Distributions in Crystals** R. S. Calder, 1961

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