

Electron Configuration of Main Group Cations

Electron configuration

Na $1s^2 2s^2 2p^6 3s^1$

Orbital diagram

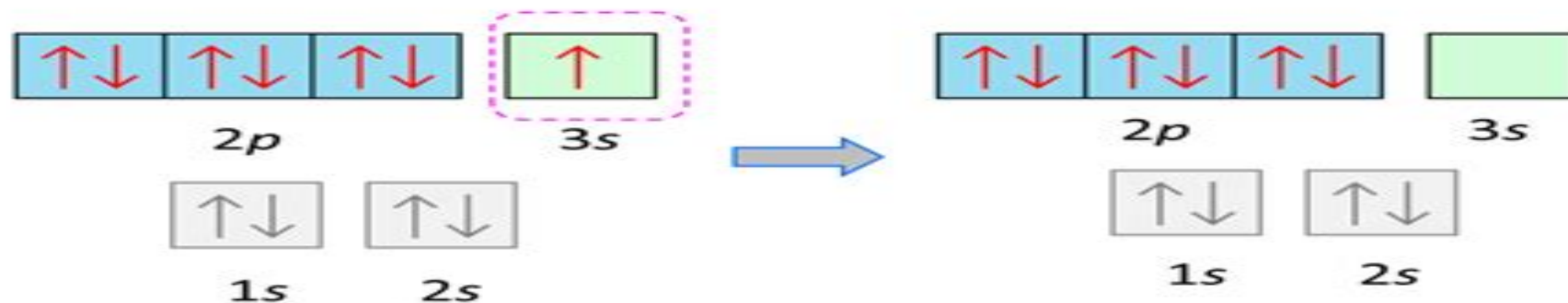


Focus on the valence shell

The charge is +1, therefore, one electron is lost:



The electron must be from the outermost valence orbital – $3s^1$



Na $1s^2 2s^2 2p^6 3s^1$

Na⁺ $1s^2 2s^2 2p^6$

↓ Same as

Ne $1s^2 2s^2 2p^6$



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Cations of main group metals attain the electron configuration of the noble gas that precedes them in the periodic table. They are **isoelectronic** with the previous noble gas.

Electron Configuration Of Atoms And Ions Lab

**Uno Kask, Otto W. Nitz, Martin G.
Ondrus, Tammy J. Melton**



Electron Configuration Of Atoms And Ions Lab:

Chemistry in the Laboratory James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations Chemistry Dorin, Prentice Hall Dictionary Editors, 1992 Nuclear Science Abstracts, 1974 2024-25 NVS Lab Attendant/Assistant Solved Papers YCT Expert Team, 2024 25 NVS Lab Attendant Assistant Solved Papers 592 995 Bilingual E This book contains previous year solved papers 66 sets and 5875 objective questions *Transplutonium Elements* Charles E. Stuber, 1972 *Nuclear Science Abstracts*, 1972 **Scientific and Technical Aerospace Reports**, 1994 Energy Research Abstracts, 1984

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features a revised art programme and explains the rational for each experiment

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