

Extension Questions

24. The hydrogen spectral lines in Model 2 are only the wavelengths of light that are in the visible range and therefore "seen" by the naked eye. However, many other wavelengths can be detected with special equipment.

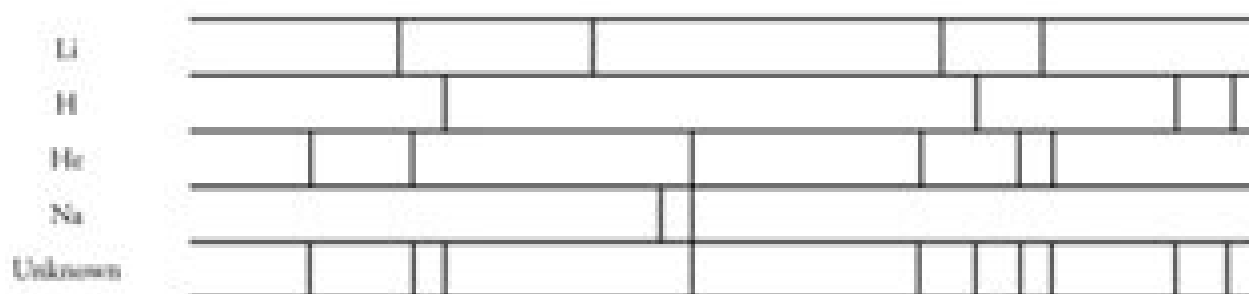
- a. Propose a hydrogen electron transition that involves light with a wavelength in the ultraviolet (UV) range (10–400 nm).

Wavelengths within the UV range would have more energy than violet wavelengths. This transition would be from $n=6$ to $n=2$ or maybe $n=6$ to $n=1$.

- b. Propose a hydrogen electron transition that involves light with a wavelength in the infrared (IR) range (1000–106 nm).

Wavelengths within the IR range would have less energy than red wavelengths. This transition would be from $n=3$ to $n=2$ or maybe $n=5$ to $n=4$.

25. Below are diagrams for the bright line spectra of four elements and the spectrum of a mixture of unknown gases.



- a. Which element(s) are not present in the Unknown?

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W. Hoffmann, D. Wegener, J. Spengler



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