# Electron Configuration Practice Worksheet KEY

In the space below, write the unabbreviated electron configurations of the following elements:

- 1) sodium 1s2 2s2 2p6 3s1
- 2) iron 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>6</sup> 4s<sup>2</sup> 3d<sup>6</sup>
- 3) bromine 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>6</sup> 4s<sup>2</sup> 3d<sup>10</sup> 4p<sup>5</sup>
- 4) barium 1s² 2s² 2p⁵ 3s² 3p⁵ 4s² 3d¹⁰ 4p⁶ 5s² 4d¹⁰ 5p⁶ 6s²
- 5) neptunium 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>6</sup> 4s<sup>2</sup> 3d<sup>10</sup> 4p<sup>6</sup> 5s<sup>2</sup> 4d<sup>10</sup> 5p<sup>6</sup> 6s<sup>2</sup> 4f<sup>14</sup> 5d<sup>10</sup> 6p<sup>6</sup> 7s<sup>2</sup> 5f<sup>5</sup>

In the space below, write the abbreviated electron configurations of the following elements:

- 6) cobalt [Ar] 4s23d7
- 7) silver [Kr] 5s24d9
- 8) tellurium [Kr] 5s24d105p4
- 9) radium [Rn] 7s<sup>2</sup>
- 10) lawrencium [Rn] 7s25f146d1

Determine what elements are denoted by the following electron configurations:

- 11) 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>4</sup> Sulfur
- 12) 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>6</sup>4s<sup>2</sup>3d<sup>10</sup>4p<sup>6</sup>5s<sup>1</sup> Rubidium
- 13) [Kr] 5s<sup>2</sup>4d<sup>10</sup>5p<sup>3</sup> Antimony
- 14) [Xe] 6s<sup>2</sup>4f<sup>14</sup>5d<sup>6</sup> Osmium
- 15) [Rn] 7s25f11 Einsteinium

Determine which of the following electron configurations are not valid:

- 16) 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3p<sup>6</sup>4s<sup>2</sup>4d<sup>10</sup>4p<sup>5</sup> should be 3d<sup>10</sup>
- 17) 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>3d<sup>5</sup> 3s<sup>2</sup> is impossible& "p" instead of "d"
- 18) [Ra] 7s<sup>2</sup>5f<sup>8</sup> Ra is not a noble gas
- 19) [Kr] 5s<sup>2</sup>4d<sup>10</sup>5p<sup>5</sup> Valid
- 20) [Xe] Element can't be its own electron configuration.

**Padhraic Smyth** 

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# Table of Contents Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key

- 1. Understanding the eBook Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
  - The Rise of Digital Reading Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
  - Personalized Recommendations
  - Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key User Reviews and Ratings
  - Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key and Bestseller Lists
- 5. Accessing Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key Free and Paid eBooks
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  - Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key eBook Subscription Services

- Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key Budget-Friendly Options
- 6. Navigating Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key eBook Formats
  - o ePub, PDF, MOBI, and More
  - Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key Compatibility with Devices
  - Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
  - Highlighting and Note-Taking Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
  - Interactive Elements Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
- 8. Staying Engaged with Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
- 9. Balancing eBooks and Physical Books Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
  - Setting Reading Goals Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
  - Fact-Checking eBook Content of Electron Configuration Worksheet Vandenbout Brake Ch301 Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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