

10. Electrolysis of copper – refining copper

Electrodes: metallic copper

Electrolyte: copper sulfate solution

Ions Present:  $\text{Cu}^{2+}$ ,  $\text{SO}_4^{2-}$ ,  $\text{OH}^-$ ,  $\text{H}^+$

Half equations:

In the Cathode  $\text{Cu}^{2+} + 2e^- \rightarrow \text{Cu}$

In the Anode  $\text{Cu} \rightarrow \text{Cu}^{2+} + 2e^-$

Overall Equation  $\text{Cu}^{2+} + \text{Cu} \rightarrow \text{Cu} + \text{Cu}^{2+}$

Observations:

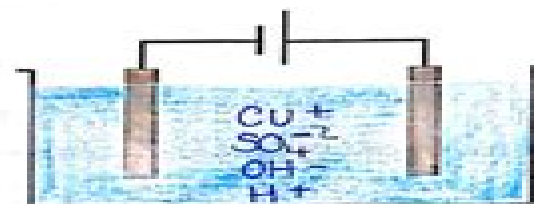
Cathode: Dark solid deposit (pure copper)

Anode: Disintegration of anode and yellow deposit below

Uses of copper: electrical wiring, copper plating

Why is it necessary to purify Copper? In this way, it will have more conductivity

What are the elements present in the slag? Ag, Au, Pt



(show in the graph how it will look like)

11. Electroplating with Copper

Electrodes:

Cathode: Spoon (metal more reactive than Cu)

Anode: Pure copper

Electrolyte: Copper Sulfate or Nitrate

Ions Present:  $\text{Cu}^{2+}$ ,  $\text{SO}_4^{2-}$ ,  $\text{OH}^-$ ,  $\text{H}^+$

Half equations:

In the Cathode  $\text{Cu}^{2+} + 2e^- \rightarrow \text{Cu}$

In the Anode  $\text{Cu} \rightarrow \text{Cu}^{2+} + 2e^-$

Overall Equation  $\text{Cu}^{2+} + 2e^- + \text{Cu} \rightarrow \text{Cu} + \text{Cu}^{2+}$

Observations:

Cathode: the color of the spoon becomes darker

Anode: Disintegrate Copper



# Electrochemistry Study Guide Answers

**AN Whitehead**



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## **Table of Contents Electrochemistry Study Guide Answers**

1. Understanding the eBook Electrochemistry Study Guide Answers
  - The Rise of Digital Reading Electrochemistry Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Electrochemistry Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Electrochemistry Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Electrochemistry Study Guide Answers

- Personalized Recommendations
- Electrochemistry Study Guide Answers User Reviews and Ratings
- Electrochemistry Study Guide Answers and Bestseller Lists
- 5. Accessing Electrochemistry Study Guide Answers Free and Paid eBooks
  - Electrochemistry Study Guide Answers Public Domain eBooks
  - Electrochemistry Study Guide Answers eBook Subscription Services
  - Electrochemistry Study Guide Answers Budget-Friendly Options
- 6. Navigating Electrochemistry Study Guide Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Electrochemistry Study Guide Answers Compatibility with Devices
  - Electrochemistry Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Electrochemistry Study Guide Answers
  - Highlighting and Note-Taking Electrochemistry Study Guide Answers
  - Interactive Elements Electrochemistry Study Guide Answers
- 8. Staying Engaged with Electrochemistry Study Guide Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Electrochemistry Study Guide Answers
- 9. Balancing eBooks and Physical Books Electrochemistry Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Electrochemistry Study Guide Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Electrochemistry Study Guide Answers
  - Setting Reading Goals Electrochemistry Study Guide Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electrochemistry Study Guide Answers

- Fact-Checking eBook Content of Electrochemistry Study Guide Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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