F332 advance notice (June 2013) "predicted Q's"

How does photolysis play a part in the destruction of N₂O in the stratosphere?

$$N_2$$
 O_{CD} + hv \longrightarrow N_2 O_{CD} + O_{CD} $O_{$

Use a series of equations to show how NO acts in the catalytic destruction of ozone?

$$\begin{array}{c} NO_2 \longrightarrow NO + O \\ NO + O_3 \longrightarrow NO_2 + O_2 \\ \hline O_3 \longrightarrow O_2 + O \end{array} \qquad \begin{array}{c} \text{This can be a boach} \\ \text{fixing at trapsopheric} \\ \text{level} \cdot NO \text{ can react} \\ \text{with } C_2 H_{Y_g} \\ \text{preventing NO} \\ \text{breaking about } O_3 \cdot \\ \text{breaking about } O_3 \cdot \\ \text{Swild up of } O_1 \text{ in That}. \\ \text{Swild up of } O_2 \text{ in That}. \\ \text{Swild up of } O_3 \cdot \\ \text{Swild up of } O_3 \cdot$$

Convert 100 parts per billion to a percentage, for the level of NOs in urban areas

1 part per million ->% (divide by 10,000)
100 ppb = 0.1 ppm (1,000 ppb = 1 ppm)

$$\frac{0.1 ppm}{10,000} = \frac{1 \times 10^{-5} \% (or, 0.00001 \%)}{}$$

Why is the NyO much into NyO may not a true dynamic equilibrium?

- Describe how photochemical smog can be harmful to: (a) the environment?
- crop damage - compounds with C=C bands are particularly vulnerable (rubbers, paints ...)

<u>June 2013 F332 Paper</u>

R Sandford

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