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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2\text{Ag}^+ \text{aq}$ $\text{Cr}_2\text{O}_7^{2-} \text{aq}$

rightarrow aq 2Cr^{2+} 7S^{2-} label 4 2 3

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web the six NO_3^{-}aq ions and the six Na^{+}aq ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+}\text{aq} + 2\text{PO}_4^{3-}\text{aq}$ $\rightarrow 3\text{Ba}_3\text{PO}_4\text{s}$

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web these spectator ions ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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web practice problems on net ionic equations page 2 of 3 answer key to practice problems on net ionic equations 1 molecular $\text{AgNO}_3\text{aq} + \text{KClaq}$ $\rightarrow \text{AgCl s} + \text{KNO}_3\text{aq}$ total ionic $\text{Ag}^{+}\text{aq} + \text{NO}_3^{-}\text{aq} + \text{K}^{+}\text{aq} + \text{Cl}^{-}\text{aq}$ $\rightarrow \text{AgCl s} + \text{K}^{+}\text{aq} + \text{NO}_3^{-}\text{aq}$ net ionic $\text{Ag}^{+}\text{aq} + \text{Cl}^{-}\text{aq}$ $\rightarrow \text{AgCl s}$ 2

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web video answer you know spectator ions are the ions present on the both sides of the equation indicate cancels out in the total ionic equations so you know in a case chemical reactions are described with chemical equations which symbolically present the starting materials end products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^{+}\text{aq} + \text{OH}^{-}\text{aq}$ $\rightarrow \text{H}_2\text{O l}$

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