

**Review for the Full Year General Chemistry
American Chemical Society Exam**

(a) Thermodynamics, (b) Transition Metals, (c) Redox and Electrochemistry, (d) Nuclear and (e) Organic

THERMODYNAMICS

84. A 10.0 g sample of silver is heated to 100.0 °C and then added to 20.0 g of water at 23.0 °C in an insulated calorimeter. At thermal equilibrium the temperature of the system was measured as 25.0 °C. What is the specific heat of silver? [specific heat = $4.2 \text{ J g}^{-1} \text{ }^{\circ}\text{C}^{-1}$]

- (A) 0.11 J/g K (C) 17 J/g K
(B) 0.22 J/g K (D) 34 J/g K

85. What is ΔH in kJ for the reaction:



- (A) -1830 kJ (B) -2284 kJ
(C) -2510 kJ (D) -1605 kJ

86. Given these equations:



calculate the heat of formation of $\text{SO}_3(\text{g})$.

- (A) -500 kJ mol⁻¹ (C) +100 kJ mol⁻¹
(B) -400 kJ mol⁻¹ (D) +200 kJ mol⁻¹

87. Given:



How much energy is produced in the combustion of 28 g of CO? [Atomic Masses: C 12.0 g mol⁻¹; O 16.0 g mol⁻¹]

- (A) 88 kJ (B) 109 kJ
(C) 175 kJ (D) 284 kJ

88. Calculate the value of ΔH (in kJ mol⁻¹) for the reaction:



Bond Energies (kJ mol⁻¹)

H-H	435
N=N	945 (in N ₂)
N-H	389

- (A) 2340 kJ of heat absorbed
(B) 213 kJ of heat absorbed
(C) 2340 kJ of heat evolved
(D) 83 kJ of heat evolved

89. Which change is likely to be accompanied by the greatest increase in entropy?



90. For which process is the entropy change per mole the largest at constant temperature?



91. In which process is entropy decreased?

- (A) dissolving sugar in water
(B) expanding a gas
(C) evaporating a liquid
(D) freezing water

92. When $\text{Al}_2\text{O}_3(\text{s})$ is formed from the elements at standard conditions, the values of ΔH° and ΔG° at 298 K are -1676 kJ mol⁻¹ and -1577 kJ mol⁻¹, respectively. The standard entropy of formation per mole, in joules per degree, will be

- (A) -332 (B) -157 (C) -93.3
(D) -0.0933 (E) +15.7

93. Vaporization of a liquid is an example of a process for which

- (A) ΔH , ΔS , and ΔG are positive at all temperatures.
(B) ΔH and ΔS are positive.
(C) ΔG is negative at low temperatures, positive at high temperatures.
(D) $\Delta H = \Delta S$

94. A particular chemical reaction has a negative ΔH and negative ΔS . Which statement is correct?

- (A) The reaction is spontaneous at all temperatures.
(B) The reaction is nonspontaneous at all temperatures.
(C) The reaction becomes spontaneous as temperature increases.
(D) The reaction becomes spontaneous as temperature decreases.

General Chemistry Exam 4

**William Crookes, James H.
Gardiner, Gerald Druce, H. W. Blood-
Ryan**

General Chemistry Exam 4:

Teach Students How to Learn Saundra Yancy McGuire, 2023-07-03 Co published with NISOD Miriam a freshman Calculus student at Louisiana State University made 37 5% on her first exam but 83% and 93% on the next two Matt a first year General Chemistry student at the University of Utah scored 65% and 55% on his first two exams and 95% on his third These are representative of thousands of students who decisively improved their grades by acting on the advice described in this book What is preventing your students from performing according to expectations Saundra McGuire offers a simple but profound answer If you teach students how to learn and give them simple straightforward strategies to use they can significantly increase their learning and performance For over a decade Saundra McGuire has been acclaimed for her presentations and workshops on metacognition and student learning because the tools and strategies she shares have enabled faculty to facilitate dramatic improvements in student learning and success This book encapsulates the model and ideas she has developed in the past fifteen years ideas that are being adopted by an increasing number of faculty with considerable effect The methods she proposes do not require restructuring courses or an inordinate amount of time to teach They can often be accomplished in a single session transforming students from memorizers and regurgitators to students who begin to think critically and take responsibility for their own learning Saundra McGuire takes the reader sequentially through the ideas and strategies that students need to understand and implement First she demonstrates how introducing students to metacognition and Bloom s Taxonomy reveals to them the importance of understanding how they learn and provides the lens through which they can view learning activities and measure their intellectual growth Next she presents a specific study system that can quickly empower students to maximize their learning Then she addresses the importance of dealing with emotion attitudes and motivation by suggesting ways to change students mindsets about ability and by providing a range of strategies to boost motivation and learning finally she offers guidance to faculty on partnering with campus learning centers She pays particular attention to academically unprepared students noting that the strategies she offers for this particular population are equally beneficial for all students While stressing that there are many ways to teach effectively and that readers can be flexible in picking and choosing among the strategies she presents Saundra McGuire offers the reader a step by step process for delivering the key messages of the book to students in as little as 50 minutes Free online supplements provide three slide sets and a sample video lecture This book is written primarily for faculty but will be equally useful for TAs tutors and learning center professionals For readers with no background in education or cognitive psychology the book avoids jargon and esoteric theory

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Wahlestedt,2016-12-19 Reviews key general chemistry concepts and techniques adapted for application to important organic principles Provides practical guidance to help students make the notoriously well known and arduous transition from general chemistry to organic chemistry Explains organic concepts and reaction mechanisms generally expanding the focus on how to understand each step from a more intuitive viewpoint Covers concepts that need further explanation as well as those that summarize and emphasize key ideas or skills necessary in this field An added bonus is help with organizing principles to make sense of a wide range of similar reactions and mechanisms Implements a user friendly process to achieve the end result of problem solving Covers organic chemistry I and II concepts at the level and depth of a standard ACS organic chemistry curriculum features practice problems and solutions to help master the material including an extensive and comprehensive bank of practice exams with solutions **Peterson's Graduate Programs in the Physical Sciences 2011**

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Chemical News and Journal of Physical Science William Crookes,James H. Gardiner,Gerald Druce,H. W. Blood-Ryan,1916

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