

Chemistry Benchmark II Study Guide

- Write the formula for the following:
 - iron (II) oxide
 - calcium phosphate
 - potassium bromide
 - ammonium phosphate
 - copper (II) bromide
 - magnesium hydroxide
 - nitrogen dioxide
 - titanium (IV) oxide
 - xenon hexafluoride
- What is a binary compound? What is a ternary compound?
- Write formulas for
 - hydroiodic acid
 - carbonic acid
 - phosphorous acid
 - sulfuric acid
 - hydrosulfuric acid
 - hydrochloric acid
 - chloric acid
 - chlorous acid
 - hypochlorous acid
 - perchloric acid
- Name these:
 - NaCl
 - KI
 - CsBr
 - Rb₂O
 - SrI₂
 - CuCl
 - HgO
 - Fe₂O₃
 - MnO₂
 - PbCl₄
 - CuI
 - SnBr₄
 - BF₃
 - NO
 - N₂O₅
 - CCl₄
 - P₂O₅
 - N₂O₃
- Find the oxidation number assigned to:
 - S in H₂SO₄
 - P in PCl₅
 - S in SO₃
- Name these.
 - NO₂
 - N₂O
 - SO₃
- How many valence electrons do elements have in the following groups?
 - alkali metals
 - alkaline earth metals
 - halogens
 - noble gases
- Write the complete electron configuration, noble gas configuration, Lewis dot structure, and orbital notation for:
 - Sr
 - Br
 - Na
 - K
 - N
 - Cl
- In each group, identify the atom with the largest atomic radius.
 - Sn, Xe, Rb, Sr
 - Rn, He, Xe, Kr
 - Pb, Ba, Cs, At
- In each group, identify the atom with the highest ionization energy.
 - Cs, K, Li
 - Ba, Sr, Ca
 - I, Br, Cl
 - Mg, Si, S
- Define the following: ionic bond, covalent bond, polar covalent bond, nonpolar covalent bond
- Define electronegativity and describe the trend across a period and down a group. What type of element has the highest electronegativity? What type has the lowest electronegativity?
- Identify the type of bond in each:
 - K-N
 - Cs-O
 - C-N
 - O-F
 - Na-Cl
 - Cl-F
 - Cl-O
 - Ca-Cl
- Who first arranged the periodic table in order of increasing atomic mass?
- Draw Lewis structures for the following and predict the molecular shape.
 - H₂S
 - PH₃
 - O₂
 - sulfate ion
 - SO₂
 - BF₃
- Calculate the wavelength and energy of 96.6MHz radio waves.
- What is the frequency and energy of electromagnetic radiation having a wavelength of $3.33 \times 10^{-6}\text{m}$?
- Predict the ions that the following elements might form:
 - F
 - K
 - I
 - Ga
 - Fe (trick question)
 - P
 - Ag

General Chemistry Benchmark 2 Study Guide

Rosina Ehmann



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developed every year a basic education is essential to address these new challenges and this work gives such training. Written with the regulatory framework in mind the material guides readers on modelling how to conduct assessments and human and wildlife risk focusing on effects on animals rather than transport of chemicals. Simple discussions of chemistry are complemented by coverage on the behavior of the animal dynamics of the ecosystem real life situations like drought and predators in the system i.e. the natural system versus the lab setting. The book's first section contains chapters on the principles of contaminant toxicology including a brief history of the science of ecotoxicology basic principles of the science testing methods and ways of determining if animals have been exposed to either acute or chronic concentrations of contaminants. The second section deals with the primary classes of contaminants including their chemical characteristics sources uses and effects on organisms. The third section focuses on more complex issues such as the regulation of pollution population and community effects risk assessment and modelling. Uses examples from both aquatic and terrestrial environments and species. Includes a Terms to Know section and a list of study questions in each chapter fostering a greater understanding of the issues. Focuses on the effects of contaminants on wildlife while providing enough chemistry to allow a detailed understanding of the various contaminant groups. Emphasizes natural examples and real species rather than laboratory studies on only a handful of organisms. Features case histories detailing actual events that include aspects of how the contamination occurred and its effects on wildlife. Provides material from a wide variety of international sources.

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