

1. According to the Brønsted—Lowry definition, which chemical species can function *both* as an acid and as a base?

- $\text{Cl}^-$
- $\text{SO}_4^{2-}$
- $\text{NH}_4^+$
- $\text{HCO}_3^-$
- $\text{H}_3\text{O}^+$

2. In the reaction  $\text{CN}^- + \text{H}_2\text{O} \rightleftharpoons \text{HCN} + \text{OH}^-$  which is an acid—base conjugate pair?

- $\text{H}_2\text{O}$  and  $\text{HCN}$
- $\text{H}_2\text{O}$  and  $\text{OH}^-$
- $\text{CN}^-$  and  $\text{H}_2\text{O}$
- $\text{HCN}$  and  $\text{OH}^-$

3. Given that HX is a stronger Bronsted acid than HY in aqueous solution, which is true of a 1 M solution of NaX

- It is less basic than a 1 M solution of NaY.
- It is more basic than a 1 M solution of NaY.
- It yields a neutral solution.
- It is more concentrated than a 1 M solution of NaY.

4. The pH of a 0.03 M HCl solution is

a.	1.5
b.	2.5
c.	3.5
d.	12.5

5. The pH of a  $1.0 \times 10^{-3}$  M  $\text{Ba}(\text{OH})_2$  solution at  $25^\circ\text{C}$  is

a.	2.7
b.	3.0
c.	11.0
d.	11.3

6. Which is the strongest acid?

- $\text{HClO}$
- $\text{HClO}_2$
- $\text{HClO}_3$
- $\text{HClO}_4$

7. Which salt reacts with water (hydrolyzes) to produce a basic solution?

a.  $\text{NaC}_2\text{H}_3\text{O}_2$   
b.  $\text{NaNO}_3$   
c.  $\text{NH}_4\text{Cl}$   
d.  $\text{BaSO}_4$

8. The weakest of the bases listed is

Acid	Conjugate Base	K <sub>a</sub> (Ionization Constant of Acid)
HCl	Cl <sup>-</sup>	100% ionized
HSO <sub>4</sub> <sup>-</sup>	SO <sub>4</sub> <sup>2-</sup>	1.2 x 10 <sup>-2</sup>
H <sub>2</sub> S	HS <sup>-</sup>	5.7 x 10 <sup>-8</sup>
HS <sup>-</sup>	S <sup>2-</sup>	1.2 x 10 <sup>-14</sup>

- $\text{Cl}^-$
- $\text{CN}^-$
- $\text{HS}^-$
- $\text{S}^{2-}$
- $\text{SO}_3^{2-}$

[illegible]

10. Which statement is a logical inference from the fact that a 0.10 M solution of potassium acetate  $\text{KC}_2\text{H}_3\text{O}_2$  is less alkaline than a 0.10 M solution of potassium cyanide, KCN?

- Hydrocyanic acid is a weaker acid than acetic acid.
- Hydrocyanic acid is less soluble in water than acetic acid.
- Cyanides are less soluble than acetates.
- Acetic acid is a weaker acid than hydrocyanic acid.

11. In the titration of 50.0 mL of 0.100 M benzoic acid (a monoprotic acid) with 50.0 mL of 0.100 M NaOH, the properties of the solution at the equivalence point will correspond exactly to the properties of

- 0.100 M sodium solution,
- 0.0500 M sodium hydroxide solution,
- 0.0500 M benzoic acid solution,
- 0.0500 M sodium benzoate solution.

# General Chemistry 1 Acs Practice Test

**Lingjun Ying**



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