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 determine based on electronegativity and symmetry if the following molecules are polar or nonpolar
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 a. cl₂ b. clo₂ c. tecl₂ d. pcl₃ e. sef₄ f. ph₂ g. xef₂ answer problem 6.2.3.6.2.3

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web the relative electronegativities of the bonded atoms is h < c < cl and so the bond moments all point toward the cl end of the

molecule and sum to yield a considerable dipole moment the molecules are relatively polar for molecules of high symmetry such as BF_3 trigonal planar CH_4 tetrahedral PF_5 trigonal bipyramidal and SF_6

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