# **CLASSIFYING CHEMICAL REACTIONS**

14. SO<sub>3</sub> + H<sub>2</sub>O → H<sub>2</sub>SO<sub>4</sub>



Classify the following reactions as synthesis, decomposition, single replacement or double replacement. (AACCI = RC)

repr	acement.	ANSW	EKS/
1.	2KCIO <sub>3</sub> → 2KCI + 3O <sub>2</sub>	Decon	nposition
2.	HCI + NaOH → NaCI + H <sub>2</sub> O	Double	Replacement
з.	Mg + 2HCl → MgCl <sub>2</sub> + H <sub>2</sub>	Single	Replacement.
4.	2H <sub>2</sub> + O <sub>2</sub> → 2H <sub>2</sub> O	Synthesis	
5.	2AI + 3NIBr₂ → 2AIBr₃ + 3NI	Single	Replacement
6.	4AI + 3O₂ → 2AI₂O₃	Synth	ne Sis
7.	2NaCl → 2Na + Cl <sub>2</sub>	Decom	position
8.	$CaCl_2 + F_2 \rightarrow CaF_2 + Cl_2$	Single	Replacement
9.	AgNO <sub>3</sub> + KCI → AgCI + KNO <sub>3</sub>	Double	Replacement
10.	N <sub>2</sub> + 3H <sub>2</sub> → 2NH <sub>3</sub>	Synth	nesis
11.	2H <sub>2</sub> O <sub>2</sub> → 2H <sub>2</sub> O + O <sub>2</sub>	Decomp	position
12.	$(NH_4)_2SO_4 + Ba(NO_3)_2 \rightarrow BaSO_4$	+ 2NH4NO3 DO	uble Replacement
13.	Mgl <sub>2</sub> + Br <sub>2</sub> → MgBr <sub>2</sub> + l <sub>2</sub>	Single	Replacement

Synthesis

# **General Chemistry Massing Activities Lab Answers**

Raffaela Di Napoli

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